Electron Configuration Practice Worksheet KEY

In the space below, write the unabbreviated electron configurations of the following elements:

- sodium 1s² 2s² 2p⁶ 3s¹
- 2) iron 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d⁶
- bromine 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d¹⁰ 4p⁶
- 4) barium 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p6 5s2 4d10 5p6 6s2
- 5) neptunium 1s² 2s²2p⁶3s²3p⁶4s² 3d⁶⁰ 4p⁶ 5s² 4d⁶⁰ 5p⁶ 6s² 4f⁶⁴ 5d⁶⁰ 6p⁶ 7s² 5f⁶

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt [Ar] 4s²3d⁷
- 7) silver [Kr] 5s²4d⁹
- 8) tellurium [Kr] 5s24d105p4
- radium [Rn] 7s²
- 10) lawrencium [Rn] 7s25f46d1

Determine what elements are denoted by the following electron configurations:

- 11) 1s²2s²2p⁶3s²3p⁴ Sulfur
- 12) 1s²2s²2p⁶3s²3p⁶4s²3d¹⁰4p⁶5s¹ Rubidium
- 13) [Kr] 5s²4d¹⁰5p³ Antimony
- 14) [Xe] 6s²4f¹⁴5d⁶ Osmium
- 15) [Rn] 7s25f11 Einsteinium

Determine which of the following electron configurations are not valid:

- 16) 1s²2s²2p⁶3s²3p⁶4s²4d¹⁰4p⁵ should be 3d¹⁰
- 17) 1s²2s²2p⁶3s¹3d⁵ 3s³ is impossible& "p" instead of "d"
- 18) [Ra] 7s²5f⁸ Ra is not a noble gas
- 19) [Kr] 5s²4d¹⁰5p⁵ Valid
- 20) [Xe] Element can't be its own electron configuration.

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R Barnett

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