

Electron Configuration Practice Worksheet **KEY**

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ **Sulfur**
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ **Rubidium**
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ **Antimony**
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ **Osmium**
- 15) $[Rn] 7s^2 5f^{11}$ **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ **should be $3d^{10}$**
- 17) $1s^2 2s^2 2p^6 3s^2 3d^5$ **$3s^2$ is impossible & "p" instead of "d"**
- 18) $[Ra] 7s^2 5f^8$ **Ra is not a noble gas**
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ **Valid**
- 20) $[Xe]$ **Element can't be its own electron configuration.**

Electron Configuration Problems And Answers

Wolfgang Guggemos



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Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm, Robert J. Salm Ph.D., 2005. This drill book contains many common problem types that are asked in General Chemistry classes in High School and College. This work will give you practice with the major problem types as you prepare for finals and standardized tests.

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