

electrochemistry - This study of electrochemical cells which are applications of redox reactions.

Standard conventions:

- Standard electrode is 1 M aqueous solutions of ions that participate in the redox reaction.
- In cases where the standard ion is not present in the cell.
- In standard conditions - standard ion is present in 1 M solution.
- Conditions are written with the cell notation cell notation.

Standard Cell notation:

Anode (oxidation)		Cathode (reduction)	
<u>Electrode</u>	<u>Electrolyte</u>	<u>Electrode</u>	<u>Electrolyte</u>
<u>Electrode</u>	<u>Electrolyte</u>	<u>Electrode</u>	<u>Electrolyte</u>
<u>Electrode</u>	<u>Electrolyte</u>	<u>Electrode</u>	<u>Electrolyte</u>
<u>Electrode</u>	<u>Electrolyte</u>	<u>Electrode</u>	<u>Electrolyte</u>
<u>Electrode</u>	<u>Electrolyte</u>	<u>Electrode</u>	<u>Electrolyte</u>

Example:



Standard Cell notation:

1. Anode

2. Cathode

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C. Remigius Fresenius



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