

Nuclear Reactions

Nuclear reactions are characterized by the following:

- They involve changes to the nucleus
 - Number of protons or neutrons in the nucleus change
- Energy changes are enormous
- Not concerned with valence electrons.
- Changes in temperature and pressure have no effect on the rate of the reaction
- Small particles are important.

Important subatomic particles

proton: ${}^1_1\text{p}$ ${}^1_1\text{H}$

neutron: ${}^1_0\text{n}$

Alpha Particle: ${}^4_2\alpha$ ${}^4_2\text{He}$ 2 protons
2 neutrons

beta particle: ${}^0_{-1}\beta$ ${}^0_{-1}\text{e}$ - high energy electrons emitted from the nucleus

positron: ${}^0_{+1}\beta$ ${}^0_{+1}\text{e}$ - positive electron: what gives a proton its positive charge

Gamma Ray: ${}^0_0\gamma$ - high energy photon

Definitions:

Isotopes - same number of protons but different mass number (mass number = # protons + # neutrons)

${}^{12}_6\text{C}$ ${}^{13}_6\text{C}$ ${}^{14}_6\text{C}$ are each an isotope of C (How many neutrons does each have?)

Nuclide (pronounced NOO clyde) - a nucleus with a specific number of protons and neutrons

Radioactive - Changes in matter resulting from SPONTANEOUS changes in the composition of the nucleus.

- Such changes are often accompanied by the creation of the subatomic particles above.
- Changes to the nucleus often result in the conversion of one element into another.

Radionuclide - nuclides that are radioactive.



Guide Nuclear Chemistry Answers

James E. Brady



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Guide to U. S. Government Publications**, 1980

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Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts. After a brief review of each section's critical ideas, students are taken through stepped-out worked examples, try-it-yourself examples, and chapter quizzes, all structured to reinforce chapter objectives and build problem-solving techniques. The solutions manual includes detailed solutions to all odd-numbered exercises in the text. **Student Solutions Manual** Darrell Ebbing, 2005-03-17 Provides solutions to odd-numbered Practice Problems, General Problems, and Cumulative Skills Problems, plus answers to Review Questions. *CCEA A2 Unit 2 Chemistry Student Guide: Analytical, Transition Metals, Electrochemistry and Organic Nitrogen Chemistry* Alyn G. McFarland, 2017-08-07 Exam Board CCEA Level A level Subject

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