



CHEMISTRY

CHEAT SHEET



NAMING SIMPLE COMPOUNDS

NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Their Formulas
(Show the work for each compound)



Carbon Dioxide (Naming Practice)

Directions for naming:

1. Name the cation first, then the anion.
2. If the cation is a transition metal, include its oxidation state in parentheses.
3. If the anion is a polyatomic ion, use its name.
4. If the cation is a transition metal, use the Roman numeral for its oxidation state.
5. If the anion is a polyatomic ion, use its name.
6. If the compound is a binary ionic compound, use the element names.
7. If the compound is a molecular compound, use the prefixes.

Binary Compounds

First Element	Anion
H ₂	hydrogen
Li	lithium
Na	sodium
K	potassium
Ag	silver
Ca	calcium
Sc	scandium
Ti	titanium
V	vanadium
Cr	chromium
Mn	manganese
Fe	iron
Ni	nickel
Cu	copper
Zn	zinc
Al	aluminum
Ga	gallium
In	indium
Sn	tin
Pb	lead
Bi	bismuth
Po	polonium

Common Ions (Polyatomic)

Formula	Name
NO ₃ ⁻	nitrate
NO ₂ ⁻	nitrite
CO ₃ ²⁻	carbonate
SO ₄ ²⁻	sulfate
SO ₃ ²⁻	sulfite
PO ₄ ³⁻	phosphate
PO ₃ ³⁻	phosphite
SiO ₄ ⁴⁻	silicate
BO ₃ ³⁻	borate
AlO ₂ ⁻	aluminate
FeO ₄ ²⁻	ferrate
CrO ₄ ²⁻	chromate
CrO ₂ ²⁻	chromite
MnO ₄ ⁻	manganate
MnO ₂ ²⁻	manganite
VO ₄ ³⁻	vanadate
VO ₂ ²⁻	vanadite
AsO ₄ ³⁻	arsenate
AsO ₂ ³⁻	arsenite
SbO ₄ ³⁻	antimonate
SbO ₂ ³⁻	antimonite
BiO ₄ ³⁻	bismutate
BiO ₂ ³⁻	bismutite

Naming Compounds with a Central and Peripheral



Carbon Dioxide

Directions for naming:

1. Name the cation first, then the anion.
2. If the cation is a transition metal, include its oxidation state in parentheses.
3. If the anion is a polyatomic ion, use its name.
4. If the cation is a transition metal, use the Roman numeral for its oxidation state.
5. If the anion is a polyatomic ion, use its name.
6. If the compound is a binary ionic compound, use the element names.
7. If the compound is a molecular compound, use the prefixes.

Common Ions

Formula	Name	Formula	Name
NO ₃ ⁻	nitrate	NO ₂ ⁻	nitrite
CO ₃ ²⁻	carbonate	SO ₄ ²⁻	sulfate
SO ₃ ²⁻	sulfite	PO ₄ ³⁻	phosphate
PO ₃ ³⁻	phosphite	SiO ₄ ⁴⁻	silicate
BO ₃ ³⁻	borate	AlO ₂ ⁻	aluminate
FeO ₄ ²⁻	ferrate	CrO ₄ ²⁻	chromate
CrO ₂ ²⁻	chromite	MnO ₄ ⁻	manganate
MnO ₂ ²⁻	manganite	VO ₄ ³⁻	vanadate
VO ₂ ²⁻	vanadite	AsO ₄ ³⁻	arsenate
AsO ₂ ³⁻	arsenite	SbO ₄ ³⁻	antimonate
SbO ₂ ³⁻	antimonite	BiO ₄ ³⁻	bismutate
BiO ₂ ³⁻	bismutite		

Common Ions for Specific Elements

Formula	Name	Formula	Name
NO ₃ ⁻	nitrate	NO ₂ ⁻	nitrite
CO ₃ ²⁻	carbonate	SO ₄ ²⁻	sulfate
SO ₃ ²⁻	sulfite	PO ₄ ³⁻	phosphate
PO ₃ ³⁻	phosphite	SiO ₄ ⁴⁻	silicate
BO ₃ ³⁻	borate	AlO ₂ ⁻	aluminate
FeO ₄ ²⁻	ferrate	CrO ₄ ²⁻	chromate
CrO ₂ ²⁻	chromite	MnO ₄ ⁻	manganate
MnO ₂ ²⁻	manganite	VO ₄ ³⁻	vanadate
VO ₂ ²⁻	vanadite	AsO ₄ ³⁻	arsenate
AsO ₂ ³⁻	arsenite	SbO ₄ ³⁻	antimonate
SbO ₂ ³⁻	antimonite	BiO ₄ ³⁻	bismutate
BiO ₂ ³⁻	bismutite		

Naming Compounds Study Guide

Robert J. Ouellette, J. David Rawn



Naming Compounds Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises [Study Guide to Accompany Basics for Chemistry](#) Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Organic Chemistry, Student Study Guide and Solutions Manual** David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become

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Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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