

Introduction:

Water is an integral part of many ionic solids and such ionic solids are called *hydrates*. The water in these solids is called *waters of hydration*. A familiar example of a hydrate is plaster of Paris, which is the hemihydrate of calcium sulfate, $\text{CaSO}_4 \bullet \frac{1}{2} \text{H}_2\text{O}$. When water is added to plaster of Paris and the material is allowed to set, it is gradually transformed into a hard crystalline compound, calcium sulfate dihydrate, $\text{CaSO}_4 \bullet 2\text{H}_2\text{O}$. This is the material of plaster casts. The difference in composition between plaster of Paris and the plaster in casts is directly associated with the different degree of hydration of the calcium sulfate in the two cases.

The water of hydration is not as tightly bound in the hydrated crystal as the ions are. The water can usually be driven off by heating the crystals in a burner flame. The material that remains after the water has been removed is called the anhydrous salt.



An example of such a change is seen in the heating of calcium sulfate dihydrate.



In this experiment, you will heat several hydrates and observe the appearance of the reactants and products involved. A non-hydrated compound will also be heated to serve as a basis for comparison.

Objectives

1. Observe the effect of heat on hydrated ionic salts.
2. Verify the hydrate formula using stoichiometric methods

Procedure.

1. Obtain 2 test tubes and record the mass of each. Record
2. Place 2-3 grams of each substance into the appropriate test tube, mass the test tube and hydrate. Record
3. Using a test tube holder, holding the test tube nearly horizontal, gently heat the test tube. Note any changes in the appearance in the crystals.
4. Allow the test tube to cool and mass. Record
5. Repeat steps 3 & 4 until the mass of the ionic salts remains the same.
6. Repeat steps 2 -5 for the same ionic salt.

Data

Test Tube	Mass of Test Tube	Mass of Test Tube & Hydrate	Mass of Hydrate	Mass of Anhydrous	Mass of Water
1					
2					

Observations

Tube	Ionic Salt/hydrate	Observations
1		
2		

Hydrates Lab Answers

KJ Lindholm-Leary

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

Hydrates Lab Answers:

Natural Gas Hydrates in Flow Assurance Carolyn Ann Koh, Amadeu Sum, 2010-10-12 With millions of kilometres of onshore and offshore oil and gas pipelines in service around the world pipelines are the life s blood of the world Notorious for disrupting natural gas production or transmission the formation of natural gas hydrates can cost a company hundreds of millions and lead to catastrophic equipment breakdowns and safety and health hazards Written by an international group of experts Natural Gas Hydrates in Flow Assurance provide an expert overview of the practice and theory in natural gas hydrates with applications primarily in flow assurance Compact and easy to use the book provides readers with a wealth of materials which include the key lessons learned in the industry over the last 20 years Packed with field case studies the book is designed to provide hands on training and practice in calculating hydrate phase equilibria and plug dissociation In addition readers receive executable programs to calculate hydrate thermodynamics Case studies of hydrates in flow assurance The key concepts underlying the practical applications An overview of the state of the art flow assurance industrial developments

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