



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat absorbed per mole at 25°C, positive temperature change

12. Which of the following is a statement of the first law of thermodynamics?

- A. the negative change in enthalpy corresponds to an exothermic process
- B. energy lost by the system must be gained by the surroundings, neither of which are created
- C. temperature is a measure of average kinetic energy
- D. enthalpy is a state function

13. The specific enthalpy of combustion of glucose, ΔH is negative. This means that the product is _____

- A. a hydrocarbon with the same enthalpy as the reactant
- B. a hydrocarbon and the temperature of the surroundings will decrease
- C. a hydrocarbon and the surroundings will increase
- D. a hydrocarbon and the temperature of the surroundings will increase
- E. none of these is a true statement

14. The reaction



is _____ and therefore heat is _____ from the system.

- A. endothermic, released
- B. exothermic, absorbed
- C. endothermic, released
- D. exothermic, absorbed
- E. thermodynamic neither released nor absorbed

15. Heat evolved when 1 mole of glucose is oxidized when the ΔH of the reaction is -1366.8 kJ is _____



- A. -1366.8 kJ
- B. 1366.8 kJ
- C. 136.68 kJ
- D. 1366.8 kJ
- E. 13.668 kJ



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Lei Shi



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1. Understanding the eBook Mcps Honors Chemistry Exam Guide
 - The Rise of Digital Reading Mcps Honors Chemistry Exam Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcps Honors Chemistry Exam Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcps Honors Chemistry Exam Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcps Honors Chemistry Exam Guide
 - Personalized Recommendations

- Mcps Honors Chemistry Exam Guide User Reviews and Ratings
- Mcps Honors Chemistry Exam Guide and Bestseller Lists
- 5. Accessing Mcps Honors Chemistry Exam Guide Free and Paid eBooks
 - Mcps Honors Chemistry Exam Guide Public Domain eBooks
 - Mcps Honors Chemistry Exam Guide eBook Subscription Services
 - Mcps Honors Chemistry Exam Guide Budget-Friendly Options
- 6. Navigating Mcps Honors Chemistry Exam Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcps Honors Chemistry Exam Guide Compatibility with Devices
 - Mcps Honors Chemistry Exam Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcps Honors Chemistry Exam Guide
 - Highlighting and Note-Taking Mcps Honors Chemistry Exam Guide
 - Interactive Elements Mcps Honors Chemistry Exam Guide
- 8. Staying Engaged with Mcps Honors Chemistry Exam Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcps Honors Chemistry Exam Guide
- 9. Balancing eBooks and Physical Books Mcps Honors Chemistry Exam Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcps Honors Chemistry Exam Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcps Honors Chemistry Exam Guide
 - Setting Reading Goals Mcps Honors Chemistry Exam Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcps Honors Chemistry Exam Guide
 - Fact-Checking eBook Content of Mcps Honors Chemistry Exam Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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