

Experiment 6B: Calculations with a Chemical Reaction

Objectives:

1. To understand the theory of stoichiometry, including the concept of limiting reagents and percent yield, for homogeneous reaction systems.
2. To calculate the limiting reagent, assuming the limiting reagents present in the solution.
3. To calculate stoichiometrically the amount of product that is formed.
4. To calculate the theoretical maximum amount of product that could be produced.
5. To compare the theoretical amount of the actual amount of product, experimentally calculated yield.

Materials: Please refer to Experiment 6B page 95 of the Pharmacokinetics Laboratory.

Equipment:

Procedures:

1. Prepare very carefully experiment with the people's experiment.
2. Prepare a solution of the limiting reagent in a graduated cylinder.
3. Add the 10 mL of the limiting reagent solution into a graduated cylinder.
4. Repeat process for the required 10 mL. Then the graduated cylinder is used.
5. Use the 10 mL solution from the graduated cylinder for the reaction.
6. Repeat the 10 mL solution and repeat the process.
7. After the reaction is complete, use the limiting reagent solution. Then the reaction.
8. Use the graduated cylinder for the reaction, then the graduated cylinder is used.
9. Repeat the process of the limiting reagent (10 mL) and the limiting reagent.
10. Repeat the process of the limiting reagent (10 mL) and the limiting reagent.

Lab 6b Calculations With A Chemical Reaction Answers

Sebastian Brünink



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