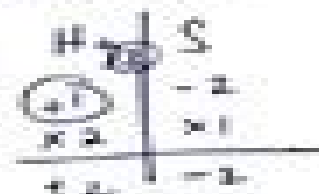
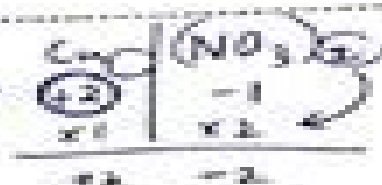


Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

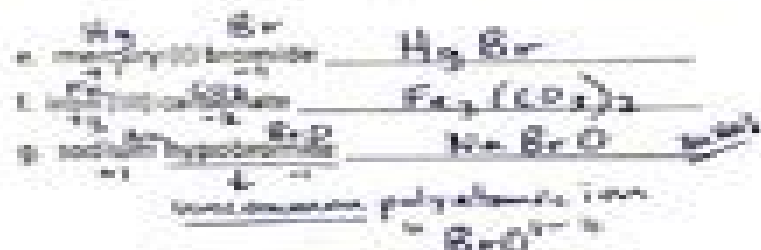
- AlF_3 Aluminum Fluoride
- Fe(OH)_2 Iron(II) Hydroxide
- $\text{Cu(NO}_3)_2$ Copper (II) Nitrate
- $\text{Ba(ClO}_3)_2$ Barium Perchlorate
- Li_3PO_4 Lithium Phosphate
- HgS Mercury (I) Sulfide
- $\text{Ca(CH}_3\text{COO)}_2$ Calcium Acetate
- $\text{Cr}_2(\text{CO}_3)_3$ Chromium (III) Carbonate
- $\text{K}_2\text{Cr}_2\text{O}_7$ Potassium Dichromate
- $\text{(NH}_4)_2\text{SO}_4$ Ammonium Sulfate



all of the underlined compounds are polyatomic ions → Memorize Them

2. Write the chemical formula for each of the following compounds:

- Copper(II) oxide Cu_2O
- potassium peroxide K_2O_2
- aluminum hydroxide Al(OH)_3
- Iron(III) bromide FeBr_3



3. Give the name or chemical formula, as appropriate, for each of the following acids:

- HNO_3 Nitric Acid
- HNO_2 Nitrous Acid
- H_3PO_4 Phosphoric Acid
phosphate PO_4^{3-} (the = oxy)

- hypochlorous acid HClO
 - iodic acid HIO_3
 - sulfurous acid H_2SO_3
- Band off of Iodate as IO_3^-

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- SiF_4 Silicon tetrafluoride
- IF_5 Iodine pentafluoride
- XeO_3 Xenon trioxide

- dinitrogen tetroxide N_2O_4
- hydrogen cyanide HCN
- tetraphosphorus hexasulfide P_4S_6

There are
all binary
ionic
compounds
(+/-)

3

Binary

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8 - 10
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13 - 14
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39 - 40
41 - 42
43 - 44
45 - 46
47 - 48
49 - 50
51 - 52
53 - 54
55 - 56
57 - 58
59 - 60
61 - 62
63 - 64
65 - 66
67 - 68
69 - 70
71 - 72
73 - 74
75 - 76
77 - 78
79 - 80
81 - 82
83 - 84
85 - 86
87 - 88
89 - 90
91 - 92
93 - 94
95 - 96
97 - 98
99 - 100

3

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1. Understanding the eBook Heath Chemistry Learning Guide Packet Answers
 - The Rise of Digital Reading Heath Chemistry Learning Guide Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Heath Chemistry Learning Guide Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heath Chemistry Learning Guide Packet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Heath Chemistry Learning Guide Packet Answers
 - Personalized Recommendations
 - Heath Chemistry Learning Guide Packet Answers User Reviews and Ratings
 - Heath Chemistry Learning Guide Packet Answers and Bestseller Lists
5. Accessing Heath Chemistry Learning Guide Packet Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Heath Chemistry Learning Guide Packet Answers
 - Interactive Elements Heath Chemistry Learning Guide Packet Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heath Chemistry Learning Guide Packet Answers
9. Balancing eBooks and Physical Books Heath Chemistry Learning Guide Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heath Chemistry Learning Guide Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Heath Chemistry Learning Guide Packet Answers
 - Setting Reading Goals Heath Chemistry Learning Guide Packet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Heath Chemistry Learning Guide Packet Answers
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 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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