

HEAT AND ITS MEASUREMENT

Name _____

Heat (or energy) can be measured in units of calories or joules. When there is a temperature change (ΔT), heat (Q) can be calculated using this formula:

$$Q = \text{mass} \times \Delta T \times \text{specific heat capacity}$$

($\Delta T = \text{Final Temp} - \text{Initial Temp}$)

During a phase change, we use this formula:

$$Q = \text{mass} \times \text{heat of fusion (or heat of vaporization)}$$

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cool from 75° C to 25° C? (Specific heat of water = 4.18 J/g° C)

$$(5)(50)(4.18) = 1045 \text{ rounded } 1000 \text{ J}$$

2. How many calories are given off by the water in Problem 1? (Specific heat of water = 1.0 cal/g° C)

$$(5)(50)(1) = 250 \text{ cal}$$

3. How many joules does it take to melt 35 g of ice at 0° C? (heat of fusion = 333 J/g)

$$(35g)(333) = 11655$$

No ΔT when melting Just 0°

12000 J

4. How many calories are given off when 85 g of steam condense to liquid water? (heat of vaporization = 539.4 cal/g)

$$46000 \text{ cal}$$

5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10° C to 60° C?

$$3000 \text{ J}$$

6. How many calories are given off when 50 g of water at 0° freezes? (heat of fusion = 79.72 cal/g)

$$4000 \text{ cal}$$

Heat And Its Measurements Chemistry 1f8766

David Ralph Dyck



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Heat and Its Uses Irving Adler, Ruth Adler, 1964 An introduction to heat covering its measurement creation conduction uses and the movement of molecules *Temperature* American Institute of Physics, 1941 *Laboratory Notes on Heat Measurements*, 1902 Science of Heat and Thermophysical Studies Jaroslav Sestak, 2005-11-15 Science of Heat and Thermophysical Studies provides a non traditional bridging of historical philosophical societal and scientific aspects of heat with a comprehensive approach to the field of generalized thermodynamics It involves Greek philosophical views and their impact on the development of contemporary ideas Covered topics include the concept of heat thermometry and calorimetry early concepts of temperature and its gradients non equilibrium and quantum thermodynamics chemical kinetics entropy order and information thermal science applied to economy econophysics ecosystems and process dynamics or mesoscopic scales quantum diffusion importance of energy science and its influence to societal life The Nature of Heat and Its Relationship to Chemistry in the Eighteenth Century David Ralph Dyck, 1967 *A Text Book of Thermo-Chemistry and Thermodynamics (Classic Reprint)* Otto Sackur, 2018-02-08 Excerpt from A d104 Book of Thermo Chemistry and Thermodynamics If we wish to study the laws of the ow of heat we must introduce exact measures for the merely psychological conceptions of heat and temperature which we have been considering hitherto Our sensation of temperature cannot help us in this as sensations cannot be compared with one another quantitatively We cannot perform a measurement of heat or of temperature until we have found some physical property of a body which varies uniquely and continuously with its thermal condition and which can be measured in time and space by one of the ordinary scientific methods Fortunately experiment has Shown us that there are a great many physical properties of bodies which vary uniquely in this way with their thermal condition and which may all therefore be used for the measurement of heat or of temperature Nearly all the physical properties of a body are altered when it is heated One of the easiest properties to measure is the volume A body occupies more space when it is hot than when it is cold We say therefore that heat causes expansion and cold causes contraction The property of expansion on heating is common to all bodies with very few exceptions one exception is water between 0 and and is generally used as the basis of temperature measurements The expansion on heating is most marked in the case of gases and was noticed first in them as early as 100 we find this property of gases made use of by Hero of Alexandria in some ingenious experiments In the case of liquids and solids the expansion is much less noticeable in the latter case it is even somewhat difficult to determine Gases and liquids are most suitable for the measurement of temperature and for the construction of temperature measuring instruments or thermometers as they are called The gas thermometer was discovered by Liquid thermometers were introduced later Fig 1 is a diagram of Galileo s air thermometer Fig 2 of a liquid thermometer as used at the present day About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses

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