HEAT AND ITS MEASUREMENT

Name ____

Heat (or energy) can be measured in units of calories or joules. When there is a temperature change (ΔT), heat (Q) can be calculated using this formula:

 $Q = mass \times \Delta T \times specific heat capacity$ ($\Delta T = Final Temp - Initial Temp)$

During a phase change, we use this formula:

Q = mass x heat of fusion (or heat of vaporization)

Solve the following problems.

 How many joules of heat are given off when 5.0 g of water cool from 75° C to 25° C? (Specific heat of water = 4.18 j/g° C)

(5)(30)(4.18) = 1045 Round

 How many calories are given off by the water in Problem 1? (Specific heat of water = 1.0 cal/g° C)

(S)(SOXI) =

- How many joules does it take to melt 35 g of ice at 0° C? (heat of fusion = 333 J/g)

No AT when melting Just 00

- 120007
- How many calories are given off when 85 g of steam condense to liquid water?

 (heat of vaporization = 539.4 cal/g)

 (1000 calories are given off when 85 g of steam condense to liquid water?

 How many joules of heat are necessary to raise the temperature of 25 g of water from 10° C to 60° C?

- .50005
- How many calories are given off when 50 g of water at 0° freezes? (heat of fusion = 79.72 cal/g)

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Heat And Its Measurement Chemistry Answers

Lingjun Ying

 $\boldsymbol{\textbf{Heat And Its Measurement Chemistry Answers:}}$

This book delves into Heat And Its Measurement Chemistry Answers. Heat And Its Measurement Chemistry Answers is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Heat And Its Measurement Chemistry Answers, encompassing both the fundamentals and more intricate discussions.

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