

## SPECIFIC HEAT WORKSHEET

**Specific heat** is the quantity of heat required to raise the temperature of one gram of a substance by one Celsius degree.



$$Q = m c \Delta T$$

where  $Q$  = heat energy,  $m$  = mass, and  $\Delta T$  = change in temp.  
Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ .

### Directions : calculating specific heat

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{1086.75}{15.75(175-25)} = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = mC(T_f - T_i) = 10.0\text{g} (0.90\text{J/g}^\circ\text{C})(55-22) = 297 \text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{67500 \text{ J}}{(1500 \text{ g})(57-32)} = 1.8 \text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mC(T_f - T_i) = 100\text{g}(4.184\text{J/g}^\circ\text{C})(37 - 4) = 14000 \text{ J}$$

5. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{455 \text{ J}}{(25\text{g})(155-25)} = 0.14 \text{ J/g}^\circ\text{C}$$



# Heat And Temperature Student Sheet Answers

**LL Leslie**



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## **Table of Contents Heat And Temperature Student Sheet Answers**

1. Understanding the eBook Heat And Temperature Student Sheet Answers
  - The Rise of Digital Reading Heat And Temperature Student Sheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Heat And Temperature Student Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Heat And Temperature Student Sheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Heat And Temperature Student Sheet Answers
  - Personalized Recommendations
  - Heat And Temperature Student Sheet Answers User Reviews and Ratings
  - Heat And Temperature Student Sheet Answers and Bestseller Lists

5. Accessing Heat And Temperature Student Sheet Answers Free and Paid eBooks
  - Heat And Temperature Student Sheet Answers Public Domain eBooks
  - Heat And Temperature Student Sheet Answers eBook Subscription Services
  - Heat And Temperature Student Sheet Answers Budget-Friendly Options
6. Navigating Heat And Temperature Student Sheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Heat And Temperature Student Sheet Answers Compatibility with Devices
  - Heat And Temperature Student Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Heat And Temperature Student Sheet Answers
  - Highlighting and Note-Taking Heat And Temperature Student Sheet Answers
  - Interactive Elements Heat And Temperature Student Sheet Answers
8. Staying Engaged with Heat And Temperature Student Sheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Heat And Temperature Student Sheet Answers
9. Balancing eBooks and Physical Books Heat And Temperature Student Sheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Heat And Temperature Student Sheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Heat And Temperature Student Sheet Answers
  - Setting Reading Goals Heat And Temperature Student Sheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Heat And Temperature Student Sheet Answers
  - Fact-Checking eBook Content of Heat And Temperature Student Sheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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