



26. Gases W and X react in a closed, rigid vessel to form gases Y and Z according to the equation above. The initial pressure of  $W_{(g)}$  is 1.20 atm and that of  $X_{(g)}$  is 1.60 atm. No  $Y_{(g)}$  or  $Z_{(g)}$  is initially present. The experiment is carried out at constant temperature. What is the partial pressure of  $Z_{(g)}$  when the partial pressure of  $W_{(g)}$  has decreased to 1.0 atm?

- (A) 0.20 atm (B) 0.40 atm (C) 1.0 atm  
(D) 1.2 atm (E) 1.4 atm

$$\begin{array}{r} W + X \rightarrow Y + Z \\ 1.2 \quad 1.6 \quad 0 \quad 0 \\ - 1.2 \\ \hline 1.0 \end{array}$$

27. On a mountaintop, it is observed that water boils at 90 °C not at 100 °C as at sea level. This phenomenon occurs because on the mountaintop the

- (A) equilibrium water vapor pressure is higher due to the higher atmospheric pressure  
(B) equilibrium water vapor pressure is lower due to the higher atmospheric pressure  
(C) equilibrium water vapor pressure equals the atmospheric pressure at a lower temperature  
(D) water molecules have a higher average kinetic energy due to the lower atmospheric pressure  
(E) water contains a greater concentration of dissolved gases

28. A 40.0 mL sample of 0.25 M KOH is added to 60.0 mL of 0.15 M  $Ba(OH)_2$ . What is the molar concentration of  $OH^-_{(aq)}$  in the resulting solution? (Assume that the volumes are additive.)

- (A) 0.10 M (B) 0.19 M (C) 0.28 M  
(D) 0.40 M (E) 0.55 M

$$\begin{aligned} .04 \times .25 &= .01 \\ .06 \times .15 &= .009 \\ \hline .019 / .1 &= .19 \end{aligned}$$



29. A 0.03 mol sample of  $NH_4NO_3(s)$  is placed in a 1 L evacuated flask, which is then sealed and heated. The  $NH_4NO_3(s)$  decomposes completely according to the balanced equation above. The total pressure in the flask measured at 400 K is closest to which of the following? (The value of the gas constant,  $R$ , is 0.082 L atm mol<sup>-1</sup> K<sup>-1</sup>.)

- (A) 3 atm (B) 1 atm (C) 0.5 atm  
(D) 0.1 atm (E) 0.03 atm

$$(1 \times P) = .09(.082)(400)$$



$$-1,323 \text{ kJ}$$

30. For the reaction of ethylene represented above,  $\Delta H$  is -1,323 kJ. What is the value of  $\Delta H$  if the combustion produced liquid water  $H_2O(l)$ , rather than water vapor  $H_2O(g)$ ? ( $\Delta H$  for the phase change  $H_2O(g) \rightarrow H_2O(l)$  is -44 kJ mol<sup>-1</sup>.)

- (A) -1,235 kJ (B) -1,279 kJ (C) -1,323 kJ  
(D) -1,367 kJ (E) -1,411 kJ

# Mcps Chemistry Semester A Exam

**BM King**



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### **Table of Contents Mcps Chemistry Semester A Exam**

1. Understanding the eBook Mcps Chemistry Semester A Exam
  - The Rise of Digital Reading Mcps Chemistry Semester A Exam
  - Advantages of eBooks Over Traditional Books
2. Identifying Mcps Chemistry Semester A Exam
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mcps Chemistry Semester A Exam
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mcps Chemistry Semester A Exam
  - Personalized Recommendations

- Mcps Chemistry Semester A Exam User Reviews and Ratings
- Mcps Chemistry Semester A Exam and Bestseller Lists
- 5. Accessing Mcps Chemistry Semester A Exam Free and Paid eBooks
  - Mcps Chemistry Semester A Exam Public Domain eBooks
  - Mcps Chemistry Semester A Exam eBook Subscription Services
  - Mcps Chemistry Semester A Exam Budget-Friendly Options
- 6. Navigating Mcps Chemistry Semester A Exam eBook Formats
  - ePub, PDF, MOBI, and More
  - Mcps Chemistry Semester A Exam Compatibility with Devices
  - Mcps Chemistry Semester A Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mcps Chemistry Semester A Exam
  - Highlighting and Note-Taking Mcps Chemistry Semester A Exam
  - Interactive Elements Mcps Chemistry Semester A Exam
- 8. Staying Engaged with Mcps Chemistry Semester A Exam
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mcps Chemistry Semester A Exam
- 9. Balancing eBooks and Physical Books Mcps Chemistry Semester A Exam
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mcps Chemistry Semester A Exam
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mcps Chemistry Semester A Exam
  - Setting Reading Goals Mcps Chemistry Semester A Exam
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcps Chemistry Semester A Exam
  - Fact-Checking eBook Content of Mcps Chemistry Semester A Exam

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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