

# Writing $K_a$



Which of the following is the correct  $K_a$  expression for the reaction shown above?

(A)  $K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2][\text{H}_2\text{O}]}$

(B)  $K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2]}$

(C)  $K_a = \frac{[\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2]}$

(D)  $K_a = \frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]}$

(E)  $K_a = \frac{[\text{HC}_2\text{H}_3\text{O}_2]}{[\text{H}_3\text{O}^+][\text{C}_2\text{H}_3\text{O}_2^-]}$

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2]}$$

# Ka Practice Chemistry

**D Keegan**



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