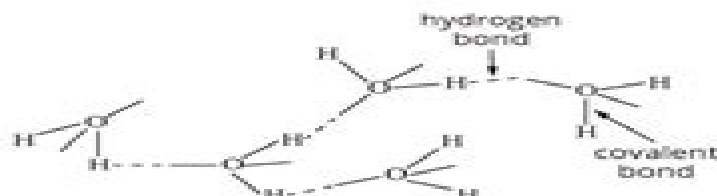


Answers to workbook worksheets

Teacher Notes

Worksheet 25 Wonderful water—structure and properties

1



- 2
- covalent
 - The O—H covalent bond is polarised because the more electronegative oxygen atom in a water molecule carries a partial negative charge and the hydrogen atoms carry a partial positive charge. The molecule is asymmetrical so it is a polar molecule.
 - see diagram for Question 1
 - Hydrogen bonding occurs between the positive partial charge on a hydrogen atom and a non-bonding pair of electrons on an oxygen atom from a different molecule.
 - examples: ammonia, hydrogen fluoride, ethanol
 - The high electronegativity of O, N and F atoms, along with their small atomic size, makes the covalent bonds they form with H atoms highly polarised. The attraction between a strong positive partial charge on the H atom and a non-bonding pair of electrons on the O, N or F atoms carrying a strong partial negative charge is relatively strong.
- 3 High latent heat of vaporisation: The water molecules can absorb a lot of energy before the hydrogen bonds between them are overcome and evaporation from liquid to gas occurs. This makes water, as sweat, a very good coolant for the human body.
- High heat capacity: Water helps to prevent overheating during exercise because the relatively strong hydrogen bonding means large amounts of energy can be absorbed by the water molecules for a relatively small increase in temperature.
- Good solvent: Nutrients for energy are delivered to the cells, and metabolic wastes are transported away and excreted, all in aqueous solution, while all of the metabolic processes for obtaining energy for exercise are aqueous reactions within cells. The ability of water to form hydrogen bonds and ion-dipole attractions to these solutes allows them to dissolve.

Worksheet 26 Specific heat capacity—calculating energy

- 1
- energy, 1°C
 - C , $\text{J g}^{-1}^{\circ}\text{C}^{-1}$, $4.18 \text{ J g}^{-1}^{\circ}\text{C}^{-1}$
 - q , $q = C \times m \times \Delta T$, specific heat capacity, mass, change in temperature
- 2
- $q = 4.18 \times 240 \times (100 - 18) = 8.23 \times 10^4 \text{ J} = 82.3 \text{ kJ}$
 - assume 3.0 L of water = 3000 g; $q = 4.18 \times 3000 \times (100 - 25) = 9.41 \times 10^5 \text{ J} = 941 \text{ kJ}$
- 3 assume 200 mL of water = 200 g; $5000 = 4.18 \times 200 \times \Delta T$, thus $\Delta T = 6.0^{\circ}\text{C}$
- 4 $1000 = 4.18 \times m \times (35 - 10)$, thus $m = 9.6 \text{ g}$

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Yi-Tong Ma

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1. Understanding the eBook Heinemann Chemistry 1 Chapter 11 Answers
 - The Rise of Digital Reading Heinemann Chemistry 1 Chapter 11 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Heinemann Chemistry 1 Chapter 11 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heinemann Chemistry 1 Chapter 11 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heinemann Chemistry 1 Chapter 11 Answers
 - Personalized Recommendations
 - Heinemann Chemistry 1 Chapter 11 Answers User Reviews and Ratings
 - Heinemann Chemistry 1 Chapter 11 Answers and Bestseller Lists
5. Accessing Heinemann Chemistry 1 Chapter 11 Answers Free and Paid eBooks
 - Heinemann Chemistry 1 Chapter 11 Answers Public Domain eBooks
 - Heinemann Chemistry 1 Chapter 11 Answers eBook Subscription Services
 - Heinemann Chemistry 1 Chapter 11 Answers Budget-Friendly Options

6. Navigating Heinemann Chemistry 1 Chapter 11 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Heinemann Chemistry 1 Chapter 11 Answers Compatibility with Devices
 - Heinemann Chemistry 1 Chapter 11 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Heinemann Chemistry 1 Chapter 11 Answers
 - Highlighting and Note-Taking Heinemann Chemistry 1 Chapter 11 Answers
 - Interactive Elements Heinemann Chemistry 1 Chapter 11 Answers
8. Staying Engaged with Heinemann Chemistry 1 Chapter 11 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heinemann Chemistry 1 Chapter 11 Answers
9. Balancing eBooks and Physical Books Heinemann Chemistry 1 Chapter 11 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heinemann Chemistry 1 Chapter 11 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Heinemann Chemistry 1 Chapter 11 Answers
 - Setting Reading Goals Heinemann Chemistry 1 Chapter 11 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Heinemann Chemistry 1 Chapter 11 Answers
 - Fact-Checking eBook Content of Heinemann Chemistry 1 Chapter 11 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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