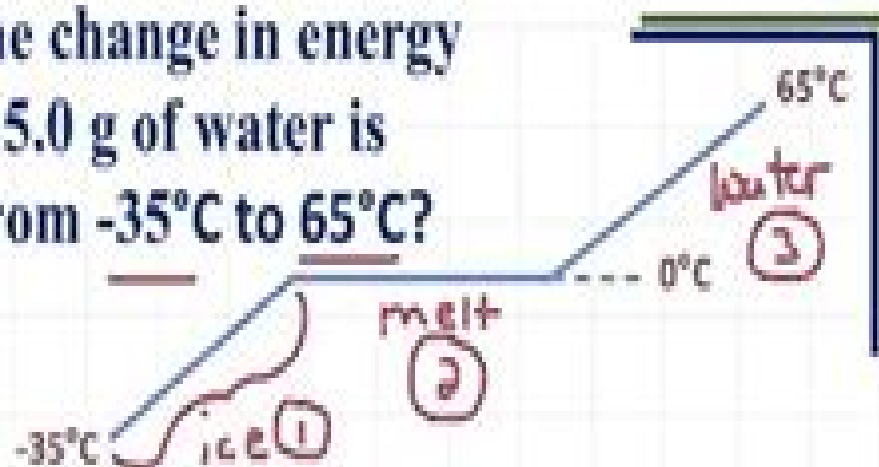


What is the change in energy  
when 25.0 g of water is  
heated from -35°C to 65°C?



$$\textcircled{1} \quad q = m C_s \Delta T = 25.00 \text{ g} \left( \frac{2.09 \text{ J}}{\text{g}^\circ\text{C}} \right) (0 - (-35))$$

$$q = 1829 \text{ J}$$

$$\textcircled{2} \quad q = n \Delta H_{\text{fus}} = 25.0 \text{ g} \left( \frac{\text{mol}}{18.01 \text{ g}} \right) \left( \frac{6.02 \text{ kJ}}{\text{mol}} \right) (1)$$

# Heating Curve Problems Key

**David Allen**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## Heating Curve Problems Key:

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