

ENTHALPY CHANGE STOICHIOMETRY ΔH of CHEMICAL REACTIONS

1) In the following reaction, what mass of propane must be burned to release 578.52 kJ of heat?
 $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g) \quad \Delta H = -2,219.2 \text{ kJ}$

2) What mass of Glucose can be produced in photosynthesis if 32763 kJ of energy are consumed in the following reaction?
 $6CO_2(g) + 6H_2O(g) \rightarrow C_6H_{12}O_6(s) + 6O_2(g) \quad \Delta H = 2838 \text{ kJ}$

3) What mass of potassium chloride can be produced when 23,834 kJ of energy are released in the following reaction:
 $C_2H_5Cl_2(l) + 2KClO_4(s) \rightarrow 2KCl(s) + 2CO_2(g) + 2H_2O(l) \quad \Delta H = -1890 \text{ kJ}$

Name _____ Section _____ Date _____

Enthalpy Change Stoichiometry

Directions: Solve each problem for the indicated quantity.

1) How much energy is absorbed when 6.827 grams of nitrogen is consumed in the following reaction?
 $N_2(g) + 2O_2(g) \rightarrow 2NO_2(g) \quad \Delta H = 66.8 \text{ kJ}$

2) How much energy is released in the formation of 25.514 grams of Cl_2 , using the following reaction?
 $2Cl(g) + O_2(g) \rightarrow 2Cl_2(g) \quad \Delta H = -983.8 \text{ kJ}$

3) Calculate the amount of energy released when acetylene burns in oxygen if 5.62 grams of acetylene are burned in the following reaction:
 $2C_2H_2(g) + 5O_2(g) \rightarrow 4CO_2(g) + 2H_2O(g) \quad \Delta H = -2599.2 \text{ kJ}$

4) How much energy is released when 13.82 grams of isooctane reacts with oxygen? Use the following reaction:
 $2C_8H_{18}(l) + 25O_2(g) \rightarrow 16CO_2(g) + 18H_2O(g) \quad \Delta H = -10912 \text{ kJ}$

5) How much energy is released when 2090 grams of NO_2 are formed in the following reaction:
 $2NO(g) + 2HCl(g) \rightarrow 2NO_2(g) + 2H_2(g) \quad \Delta H = -1009 \text{ kJ}$

6) What mass of titanium(IV) chloride is consumed in the following reaction if 583.47 kJ of energy?
 $TiCl_4(g) + 2H_2O(g) \rightarrow TiO_2(s) + 4HCl(g) \quad \Delta H = 87.2 \text{ kJ}$

7) What mass of Chlorine gas is consumed in the following reaction if 397.8 kJ are released?
 $Cl_2(g) + 2Cl(g) \rightarrow Cl_3(g) + 2Cl_2(g) \quad \Delta H = -150.3 \text{ kJ}$

A CHEMISTRY WORKSHEET

10 PROBLEMS

3 VERSIONS

WORKED ANSWER KEY

Heat Of Reaction Answer Key

Katrin Zwirgmaier

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