| | March 2010 |
|-------|------------|
| Name: | Period: |

SHOW ALL WORK FOR FULL CREDIT! Remember to include units in your answers. Round answers to the proper number of significant figures.

| Substance | Specific Heat (J/g x °C) |
|---------------|--------------------------|
| Water | 4.18 |
| Aluminum | 0.90 |
| Iron | 0.46 |
| Steam | 1.7 |
| Mercury | 0.14 |
| Chloroform | 0.96 |
| Grain Alcohol | 2.4 |
| Silver | 0.24 |
| Ice | 2.1 |

 How much heat will be absorbed by 25g of chloroform when its temperature started at 67.0°C and after it was 112.0°C?

Calculator answer:______
SF answer:

 How much heat will be absorbed by 130g of Mercury when its temperature started at 24.0°C and after it was 78.0°C?

Calculator answer:_____

 If the temperature of a 95.00 g sample of Iron started at 36.0°C and ended at 58.0°C, how much heat would be lost?

Calculator answer:_____

SF answer:_____

SF answer:

- 4. What is the definition of Temperature? _____
- 5. What is the definition of Heat? _____
- 6. Why can there never be a temperature less than 0 Kelvins? ______

Heat Capacity Practice Problems 4 Answers

DP Hallahan

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