

16.0 g of methane in an apple. Reaction contains -2812 kJ/mole.

$$\frac{16.0 \text{ g}}{(6 \times 12.01 + 2 \times 1.008 + 6 \times 5.714) \text{ g/mole}} = 8.8110904 \times 10^{-2} \text{ mole} \times -2812 \text{ kJ/mole} \times \frac{1000 \text{ J}}{1 \text{ kJ}}$$

$$= -2.48 \times 10^5 \text{ J} \times \frac{1 \text{ Cal}}{4184 \text{ J}} = -59.688244$$

150 g H_2O @ 25.1°C $C_{\text{H}_2\text{O}} = 4.18 \text{ J/g}^\circ\text{C}$ $t_f = 30.1$

21.0 g Cu @ 100.4°C $C_{\text{Cu}} = 0.385 \text{ J/g}^\circ\text{C}$ $t_f = 30.1$

$q = mc\Delta t$

$q = (150 \text{ g})(4.18 \text{ J/g}^\circ\text{C})(30.1 - 25.1)^\circ\text{C} = 3.135 \times 10^3 \text{ J} = \boxed{3.14 \times 10^3 \text{ J}}$

$q = mc\Delta t$

$q = (21.0 \text{ g})(0.385 \text{ J/g}^\circ\text{C})(100.4 - 30.1)^\circ\text{C} = \boxed{3.27 \times 10^3 \text{ J}}$

"a-b" $(3.27 \times 10^3 - 3.135 \times 10^3) \text{ J} = 1.3455 \times 10^3 \text{ J} = 1345.5 \text{ J} = \boxed{140 \text{ J}}$

Δt for calorimeter is $30.1 - 25.1 = 5^\circ\text{C} \rightarrow \frac{q}{\Delta t} = C_{\text{calor}} = \frac{1345.5 \text{ J}}{5^\circ\text{C}} = \boxed{269 \text{ J/}^\circ\text{C}}$

Heat And Thermodynamics Zemansky Solution

C. Isenberg, S. Chomet



Heat And Thermodynamics Zemansky Solution:

Unveiling the Magic of Words: A Review of "**Heat And Thermodynamics Zemansky Solution**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Heat And Thermodynamics Zemansky Solution**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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