

Thermal Energy Worksheet

1. An 80-kg patient has the flu and his temperature increases from 37°C to 39°C. How much extra heat do the chemical reactions in his cells have to generate to cause this temperature change? ($6 \times 10^5 \text{ J}$)

$$c = 3\,470 \text{ J/(kg}\cdot\text{K)}$$

2. Calculate

a) the amount of heat when 3.8 kg of ethylene glycol cools from 35°C to 20°C, $c = 2200 \text{ J/kg}\cdot\text{K}$ ($^{\circ}\text{C} = 273^{\circ}\text{K}$, therefore, $1^{\circ}\text{C} = 274^{\circ}\text{K}$) (ans: $1.3 \times 10^6 \text{ J}$)

b) the temperature change when 15 kg of mercury loses 25 kJ of heat. $c = 140 \text{ J/(kg}\cdot\text{K)}$ (ans: -12°K)

c) the mass of copper if 7.8 kJ of heat changes its temperature from 20°C to 298 K. $c = 390 \text{ J/(kg}\cdot\text{K)}$ (ans: 4.0 kg)

3. When 2.0 kg of a cold metal at a temperature of 248 K is immersed in 3.0 kg of water at a temperature of 313 K, the final temperature of the mixture is 309 K. What is the specific heat capacity of the metal? $c_w = 4200 \text{ J/(kg}\cdot\text{K)}$ (ans: $4.1 \times 10^2 \text{ J/(kg}\cdot\text{K)}$)

Use the following table to answer the questions below:

Substance	Specific Heat Capacity $\text{J/kg}\cdot^{\circ}\text{C}$	Substance	Specific Heat Capacity $\text{J/kg}\cdot^{\circ}\text{C}$
Glass	8.4×10^2	Water	4.18×10^3
Iron	4.5×10^2	Alcohol	2.5×10^3
Brass	3.8×10^2	Ice	2.1×10^3
Lead	1.3×10^2	Aluminum	9.2×10^2
Silver	2.4×10^2		

4. Calculate the amount of heat needed to raise the temperature of the following:

a) 8.4 kg of water by 6.0°C (ans: $2.1 \times 10^5 \text{ J}$)

b) 2.1 kg of alcohol by 32°C (ans: $1.7 \times 10^5 \text{ J}$)

5. Determine the heat lost when

a) 3.7 kg of water cools from 31°C to 24°C (ans: $1.1 \times 10^5 \text{ J}$)

b) a 540-g piece of silver cools from 78°C to 14°C (ans: $8.3 \times 10^3 \text{ J}$)

6. An electric immersion heater delivers 0.050 MJ of energy to 5.0 kg of a liquid, changing its temperature from 32°C to 42°C. Find the specific heat capacity of the liquid. (ans: $1.0 \times 10^3 \text{ J/(kg}\cdot^{\circ}\text{C)}$)

7. A 2.5-kg pane of glass, initially at 41°C, loses $4.2 \times 10^4 \text{ J}$ of heat. What is the new temperature of the glass? (ans: 21°C)

Heat And Heat Technology Directed Reading Answers

M Carnoy



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