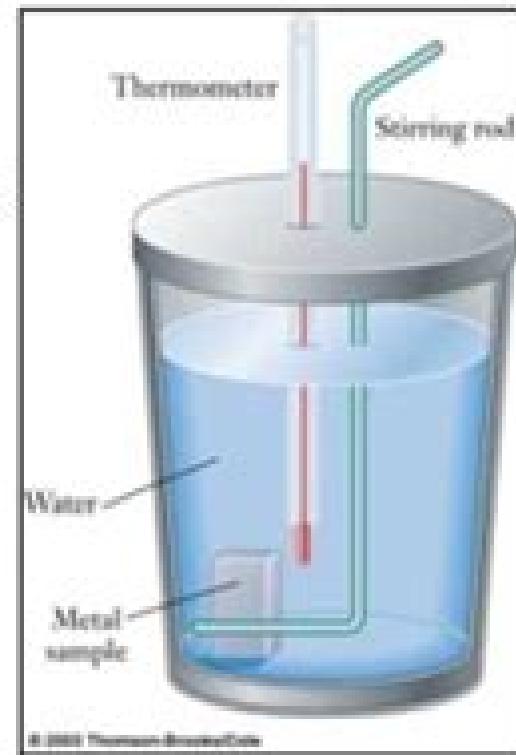


In the laboratory a "*coffee cup*" calorimeter, or constant pressure calorimeter, is frequently used to determine the specific heat of a solid, or to measure the energy of a solution phase reaction.

Since the cup itself can absorb energy, a separate experiment is needed to determine the heat capacity of the calorimeter. This is known as *calibrating* the calorimeter and the value determined is called the *calorimeter constant*.

One way to do this is to use a common metal of known heat capacity. In the laboratory a student heats 95.50 grams of zinc to 98.11 °C and then drops it into a cup containing 81.69 grams of water at 23.33 °C. She measures the final temperature to be 30.61 °C.



Using the accepted value for the specific heat of zinc (See the References tool), calculate the calorimeter constant.

Calorimeter Constant = J/°C.

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