

HEAT AND ITS MEASUREMENT

Name _____

Heat (or energy) can be measured in units of calories or joules. When there is a temperature change (ΔT), heat (Q) can be calculated using this formula:

$$Q = \text{mass} \times \Delta T \times \text{specific heat capacity}$$

$$(\Delta T = \text{Final Temp} - \text{Initial Temp})$$

During a phase change, we use this formula:

$$Q = \text{mass} \times \text{heat of fusion (or heat of vaporization)}$$

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cool from 75° C to 25° C? (Specific heat of water = 4.18 J/g° C)

$$(5)(50)(4.18) = 1045 \text{ rounded } 1000 \text{ J}$$

2. How many calories are given off by the water in Problem 1? (Specific heat of water = 1.0 cal/g° C)

$$(5)(50)(1) = 250 \text{ cal}$$

3. How many joules does it take to melt 35 g of ice at 0° C? (heat of fusion = 333 J/g)

$$(35g)(333) = 11655$$

No ΔT when melting Just 0° 12000 J

4. How many calories are given off when 85 g of steam condense to liquid water? (heat of vaporization = 539.4 cal/g)

$$46000 \text{ cal}$$

5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10° C to 60° C?

$$5000 \text{ J}$$

6. How many calories are given off when 50 g of water at 0° freezes? (heat of fusion = 79.72 cal/g)

$$4000 \text{ cal}$$

Heat And Its Measurement Chemistry If8766 Worksheet Answers

Rachel S Tattersall



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Measurement of the Specific Heat of Solids at High Temperatures Edwin Irle Griggs,1962 Heat and Temperature Measurement Robert L. Weber,1950 History of the Measurement of Heat Carl Benjamin Boyer,1943

High-resolution Specific Heat Measurements on $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ Near Its Néel Temperature ,1969 **Laboratory Notes on Heat Measurements** ,1902 **Measure to Scale! Measuring Temperature with Reading to Scale |**

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Theory of Heat James Clerk Maxwell,2020-08-22 The aim of this book is to exhibit the scientific connexion of the various steps by which our knowledge of the phenomena of heat has been extended The first of these steps is the invention of the thermometer by which the registration and comparison of temperatures is rendered possible The second step is the measurement of quantities of heat or Calorimetry The whole science of heat is founded on Thermometry and Calorimetry and when these operations are understood we may proceed to the third step which is the investigation of those relations between the thermal and the mechanical properties of substances which form the subject of Thermodynamics The whole of this part of the subject depends on the consideration of the Intrinsic Energy of a system of bodies as depending on the temperature and physical state as well as the form motion and relative position of these bodies

A Manual of Heat and Temperature Measurement ,1939 *Measuring Heat and Temperature* Arthur Spencer Callen,August Ulmann,International Correspondence Schools,1932 **Heat & Temperature Measurement** ,1982 **A Manual of measurements in heat** K. E. Guthe,1910 **The Thermochemistry of the Chemical Substances** Francis Russell Bichowsky,Frederick Dominic Rossini,1936 This reference contains a table of values with a list of substances accompanied by thermochemical data The data collected includes 5840 values of heats of formation and 350 values of heats of transition fusion vaporization and reaction All entries in the index are hyperlinked to their page numbers The Nature of Heat and Its Relationship to Chemistry in the Eighteenth Century David Ralph Dyck,1985 *The Nature of Heat and Its Relationship to Chemistry in the Eighteenth Century. [Madison, Wis.]* 1967, [c1968 David Ralph Dyck,1970 **Calorimetric Measurements of the Heat of Transport in the Diffusion of Hydrogen Through Nitrogen** Robert Joseph Zaworski,1966 Some Specific Heat Measurements at Low Temperatures John Barrington Harness,1962

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