

## CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
- Calculate the mass of FeS formed.

2. Acrylonitrile,  $\text{C}_3\text{H}_3\text{N}$ , is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene,  $\text{C}_3\text{H}_6$ , by reaction with nitric oxide,  $\text{NO}$ , as follows:



What mass of  $\text{C}_3\text{H}_3\text{N}$  can be made when 21.6 g of  $\text{C}_3\text{H}_6$  react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction:  $\text{P}_{4(s)} + 6 \text{Cl}_{2(g)} \rightarrow 4 \text{PCl}_{3(l)}$   
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid,  $\text{HCHO}_2$ , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction:  $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$
  - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl <sub>2</sub> (grams)	H <sub>2</sub> (L)
Before reaction				
After reaction	1.26 g			

# Limiting Reactant Problems Answer Key

**N Noddings**



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## **Table of Contents Limiting Reactant Problems Answer Key**

1. Understanding the eBook Limiting Reactant Problems Answer Key
  - The Rise of Digital Reading Limiting Reactant Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reactant Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Problems Answer Key
  - Personalized Recommendations
  - Limiting Reactant Problems Answer Key User Reviews and Ratings
  - Limiting Reactant Problems Answer Key and Bestseller Lists
5. Accessing Limiting Reactant Problems Answer Key Free and Paid eBooks
  - Limiting Reactant Problems Answer Key Public Domain eBooks
  - Limiting Reactant Problems Answer Key eBook Subscription Services
  - Limiting Reactant Problems Answer Key Budget-Friendly Options

6. Navigating Limiting Reactant Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reactant Problems Answer Key Compatibility with Devices
  - Limiting Reactant Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reactant Problems Answer Key
  - Highlighting and Note-Taking Limiting Reactant Problems Answer Key
  - Interactive Elements Limiting Reactant Problems Answer Key
8. Staying Engaged with Limiting Reactant Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reactant Problems Answer Key
9. Balancing eBooks and Physical Books Limiting Reactant Problems Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reactant Problems Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactant Problems Answer Key
  - Setting Reading Goals Limiting Reactant Problems Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactant Problems Answer Key
  - Fact-Checking eBook Content of Limiting Reactant Problems Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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