Problem: Reactions with Limiting Amounts of Reactants

Lithium oxide is used aboard the space shuttle to remove water from the air supply according to the equation:

$$\text{Li}_2\text{O}(s) + \text{H}_2\text{O}(g) \longrightarrow 2\text{LiOH}(s)$$

If 80.0 g of water are to be removed and 65.0 g of Li₂O are available, which reactant is limiting? How many grams of excess reactant remain? How many grams of LiOH are produced?

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