

Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

Given the following reaction: $\text{C}_5\text{H}_{12} + 8 \text{O}_2 \rightarrow 5 \text{CO}_2 + 6 \text{H}_2\text{O}$

1. If 5.25 grams of pentane (C_5H_{12}) are completely reacted with oxygen (O_2), how many grams of carbon dioxide (CO_2) will be produced?

$$5.25 \text{ g C}_5\text{H}_{12} \times \frac{1 \text{ mol C}_5\text{H}_{12}}{72.151 \text{ g C}_5\text{H}_{12}} = 0.07276 \text{ mol C}_5\text{H}_{12}$$

$$0.07276 \text{ mol C}_5\text{H}_{12} \times \frac{5 \text{ mol CO}_2}{1 \text{ mol C}_5\text{H}_{12}} = 0.3638 \text{ mol CO}_2$$

$$0.3638 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 16.0 \text{ g CO}_2$$

2. If 9.75 grams of oxygen (O_2) are completely reacted with pentane (C_5H_{12}), how many grams of carbon dioxide (CO_2) will be produced?

$$9.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 0.3047 \text{ mol O}_2$$

$$0.3047 \text{ mol O}_2 \times \frac{5 \text{ mol CO}_2}{8 \text{ mol O}_2} = 0.1904 \text{ mol CO}_2$$

$$0.1904 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 8.38 \text{ g CO}_2$$

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of CO_2 actually form?

Limiting reactant: O_2

8.38 g CO_2 actually form

Limiting Reactants Key

Bill T Reese

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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Basic Principles and Calculations in Chemical Engineering David Mautner Himmelblau, James B. Riggs, 2012 Best selling introductory chemical engineering book now updated with far more coverage of biotech nanotech and green engineering Thoroughly covers material balances gases liquids and energy balances Contains new biotech and bioengineering problems throughout

Biomass Conversion Processes for Energy and Fuels Samir S. Sofer, Oskar R. Zaborsky, 2012-12-06 Countless pages have been written on alternative energy sources since the fall of 1973 when our dependence on fossil petroleum resources became a grim reality One such alternative is the use of biomass for producing energy and liquid and gaseous fuels The term biomass generally refers to renewable organic matter generated by plants through photosynthesis Thus trees agricultural crops and aquatic plants are prime sources of biomass Furthermore as these sources of biomass are harvested and processed into commercial products residues and wastes are generated These together with municipal solid wastes not only add to the total organic raw material base that can be utilized for energy purposes but they also need to be removed for environmental reasons Biomass has been used since antiquity for energy and material needs It is still one of the most sought after energy sources in most of the world Furthermore wood was still a dominant energy source in the U S only a hundred years ago equal with coal Currently biomass contributes about 15.2 quadrillion Btu 1 quad 10¹⁵ Btu of energy to our total energy consumption of about 78 quad Two quad may not seem large when compared to the contribution made by petroleum 38 quad or natural gas 20 quad but biomass is nearly comparable to nuclear energy 2.7 quad

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