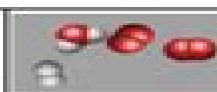


**Activity A:****Water reaction**Get the Gizmo ready:

- Make sure  $\text{H}_2 + \text{O}_2$  becomes  $\text{H}_2\text{O}$  is selected.
- Click **Reset** (↺).



**Goal:** Identify a limiting reactant

1. Count:  $\text{H}_2\text{O}$  is the **chemical formula** for water. In order to produce a single molecule of water, how many hydrogen atoms are needed?  Oxygen atoms?

2. Predict: Set the number of  $\text{O}_2$  molecules to five and the number of  $\text{H}_2$  molecules to eight.

A. How many oxygen atoms are present?  Hydrogen atoms?

B. How many water molecules could form from these reactants?

C. After the molecules react, which reactant will be left over?

D. Which reactant will be the limiting reactant?

E. Click **Play** and wait until **Reaction complete** is shown. What happened?

Due to the hydrogen and oxygen reacting, there are eight water molecules and one remaining oxygen molecule.

3. Test: Click **Reset**. Set the number of  $\text{O}_2$  molecules to 10.

A. How many hydrogen molecules ( $\text{H}_2$ ) will be needed for there to be no excess

reactants?  
20  
hydrogen  
molecules

Use the slider to set the number of  $\text{H}_2$  molecules to this value.

B. How many water molecules will be formed?

20

C. Click **Play**. What happened?

20 water molecules were formulated

4. Challenge yourself: Run the Gizmo with five  $\text{O}_2$  molecules and nine  $\text{H}_2$  molecules.

A. How many of each molecule do you think will be produced by this reaction?

$\text{O}_2$ :

$\text{H}_2$ :

$\text{H}_2\text{O}$ :

B. Click **Play**. How many were actually produced?  $\text{O}_2$ :

$\text{H}_2$ :

$\text{H}_2\text{O}$ :

# Limiting Reactant Gizmo Answer

**Ying Liu**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Limiting Reactant Gizmo Answer**

1. Understanding the eBook Limiting Reactant Gizmo Answer
  - The Rise of Digital Reading Limiting Reactant Gizmo Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Gizmo Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reactant Gizmo Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Gizmo Answer

- Personalized Recommendations
  - Limiting Reactant Gizmo Answer User Reviews and Ratings
  - Limiting Reactant Gizmo Answer and Bestseller Lists
5. Accessing Limiting Reactant Gizmo Answer Free and Paid eBooks
    - Limiting Reactant Gizmo Answer Public Domain eBooks
    - Limiting Reactant Gizmo Answer eBook Subscription Services
    - Limiting Reactant Gizmo Answer Budget-Friendly Options
  6. Navigating Limiting Reactant Gizmo Answer eBook Formats
    - ePub, PDF, MOBI, and More
    - Limiting Reactant Gizmo Answer Compatibility with Devices
    - Limiting Reactant Gizmo Answer Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Limiting Reactant Gizmo Answer
    - Highlighting and Note-Taking Limiting Reactant Gizmo Answer
    - Interactive Elements Limiting Reactant Gizmo Answer
  8. Staying Engaged with Limiting Reactant Gizmo Answer
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Limiting Reactant Gizmo Answer
  9. Balancing eBooks and Physical Books Limiting Reactant Gizmo Answer
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Limiting Reactant Gizmo Answer
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Limiting Reactant Gizmo Answer
    - Setting Reading Goals Limiting Reactant Gizmo Answer
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Limiting Reactant Gizmo Answer

- Fact-Checking eBook Content of Limiting Reactant Gizmo Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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