

## GENERAL INORGANIC CHEMISTRY

### Study of Matter

#### A. Classification of Matter

##### 1. Pure Substance

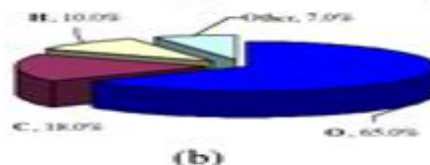
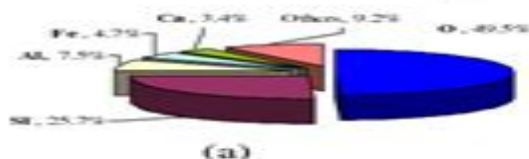
- form of matter that has distinct composition and properties

##### a. Elements

- Substance composed of one kind of atom and cannot be decomposed into simpler substances
- There are 118 known elements of which 91 elements exist in the Earth in macroscopic or trace quantities, 3 elements (Tc, Pm and Np) exist as a result of rare decay of heavy elements and 24 elements (Am to Uuo) are synthetically prepared

##### b. Compounds

- Substance formed when elements interact with other elements
- A pure compound is made up of the elements in the same proportion by mass – Law of Definite Proportions by Joseph Louis Proust (1754-1826)
- Abundance of the elements in (a) earth's crust and (b) human body



##### 2. Mixture

- form of matter that consists of different substances and whose composition varies
- a. **Heterogeneous mixture** – mixture that does not have uniform composition, properties and appearance
- b. **Homogeneous mixture** – mixture that is uniform throughout

#### B. Properties of Matter

1. **Physical properties** – observed without changing the identity and composition of the substance
2. **Chemical properties** – observed when a substance react to form other substances
3. **Intensive properties** – properties independent of the amount of substance present such as melting point, density, temperature, etc.
4. **Extensive properties** – properties dependent of the amount of substance present such as mass and volume

#### C. Changes in Matter

1. **Physical change** – changes that does not result to formation of new substances
2. **Chemical change** – changes that result to formation of new substances

### Atomic Theory

#### A. Dalton's Atomic Theory

- Proposed by John Dalton (1766-1844)

# Inorganic Chemistry Study Guide

**R. B. Heslop, Percy Lucock Robinson**



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## **Table of Contents Inorganic Chemistry Study Guide**

1. Understanding the eBook Inorganic Chemistry Study Guide
  - The Rise of Digital Reading Inorganic Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Inorganic Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Inorganic Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Inorganic Chemistry Study Guide
  - Personalized Recommendations
  - Inorganic Chemistry Study Guide User Reviews and Ratings
  - Inorganic Chemistry Study Guide and Bestseller Lists

5. Accessing Inorganic Chemistry Study Guide Free and Paid eBooks
  - Inorganic Chemistry Study Guide Public Domain eBooks
  - Inorganic Chemistry Study Guide eBook Subscription Services
  - Inorganic Chemistry Study Guide Budget-Friendly Options
6. Navigating Inorganic Chemistry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Inorganic Chemistry Study Guide Compatibility with Devices
  - Inorganic Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Inorganic Chemistry Study Guide
  - Highlighting and Note-Taking Inorganic Chemistry Study Guide
  - Interactive Elements Inorganic Chemistry Study Guide
8. Staying Engaged with Inorganic Chemistry Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Inorganic Chemistry Study Guide
9. Balancing eBooks and Physical Books Inorganic Chemistry Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Inorganic Chemistry Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Inorganic Chemistry Study Guide
  - Setting Reading Goals Inorganic Chemistry Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Inorganic Chemistry Study Guide
  - Fact-Checking eBook Content of Inorganic Chemistry Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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