

6. When one gram acid of  $\text{KMnO}_4$  reacts with  $\text{HCl}$ , the volume of chlorine liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres  
(c) 44.8 litres (d) 56.8 litres
7. When  $\text{MnO}_4^-$  ions reacts with  $\text{Fe}^{2+}$  ions in acid solution the  $\text{Fe}^{3+}$  is liberated. The equivalent weight of  $\text{KMnO}_4$  in this reaction is
- (a) 508 (b) 158  
(c) 308 (d) 348
8. 54 g. of hydrogen peroxide is present in 100 ml. of solution. This solution is called
- (a) 10 vol solution (b) 20 vol solution  
(c) 30 vol solution (d) 50 vol solution
9. The number of moles of  $\text{KMnO}_4$  that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 10 (b) 20  
(c) 40 (d) 2
10. To prepare a solution that is 0.05 M  $\text{KCl}$  starting with 100 mL of 0.40 M  $\text{KCl}$
- (a) add 0.30 g  $\text{KCl}$   
(b) add 50 mL of water  
(c) add 0.30 mol  $\text{KCl}$   
(d) evaporate 10 mL water
11. In hot acidic solution,  $\text{Br}_2$  disproportionates to  $\text{Br}^-$  and  $\text{BrO}_3^-$
- $$3\text{Br}_2 + 6\text{H}^+ \rightarrow 2\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$$
- Atomic equivalent weight of  $\text{Br}_2$  is (molecular weight = 160)
- (a)  $\frac{16}{3}$  (b)  $\frac{16}{2}$   
(c)  $\frac{160}{3}$  (d)  $\frac{160}{2}$
12.  $(\text{Fe}^{2+})$  in a solution prepared by adding 20.00 mL of 0.10 M  $\text{FeCl}_2$  with 70 mL of 0.10 M  $\text{Na}_2\text{FeO}_4$  is
- (a) 0.020 M (b) 0.140 M  
(c) 0.010 M (d) 0.200 M
13. When 80 mL of 0.100 M  $\text{HCl}$  is mixed with 120 mL of 0.10 M  $\text{KOH}$ , the resultant solution is the same as a solution of
- (a) 0.10 M  $\text{HCl}$  and 0.02 M  $\text{HCl}$   
(b) 0.08 M  $\text{HCl}$   
(c) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{KOH}$   
(d) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{HCl}$
14. 50 mL of  $\text{H}_2\text{O}_2$  solution contains strength = x% required 10 mL of  $\text{KMnO}_4$  solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.4  
(c) 0.5 (d) 06.8
15. Molarity of 1.0 M  $\text{H}_2\text{SO}_4$  (d = 1.8 mL/cm<sup>3</sup>) is
- (a) 36 mol kg<sup>-1</sup> (b) 300 mol kg<sup>-1</sup>

# Inorganic Chemistry Multiple Choice Questions With Answers

**H Kauffman**



## **Inorganic Chemistry Multiple Choice Questions With Answers:**

## Reviewing **Inorganic Chemistry Multiple Choice Questions With Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Inorganic Chemistry Multiple Choice Questions With Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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