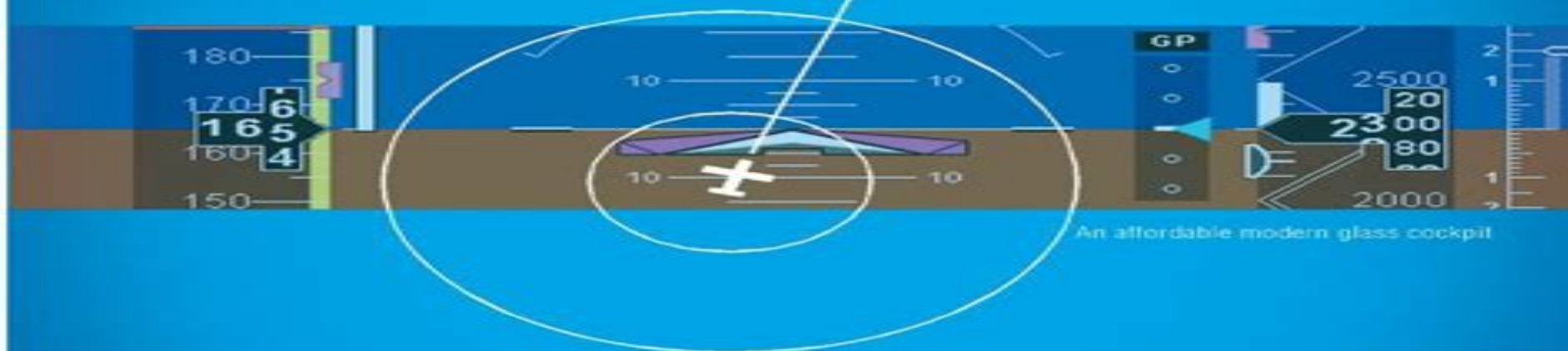


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web answers 1 hc 2h 3o 2 l 2o 2 g co 2 g 2h 2o l hc 2h 3o 2 l 4o 2 g 871 7 kj 2co 2 g 2h 2o l 2 the enthalpy of reaction is h prod h react the reactant is at a higher potential than the product when the reaction occurs heat is released stabilizing the system and the system moves to a lower energy potential 3

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in