

25. Consider the following reaction equation:



Increasing the temperature of the reaction would

- A. shift the equilibrium to the right.
 B. increase the yield of ammonia.
 C. decrease the amount of hydrogen.
 D. decrease the yield of ammonia.
26. When air in a syringe is compressed such that there is no change in temperature, the
- A. air liquefies.
 B. pressure increases.
 C. intermolecular space increases.
 D. density decreases.
27. Which of the following statements about liquids is/are true?
- I. Liquids maintain their volume at constant temperature.
 II. Liquids have fixed shape.
 III. Liquids do not diffuse.
 IV. Change in pressure affects volume of liquids.
- A. I only
 B. IV only
 C. I and IV only
 D. II and III only
28. A hydrogen chloride gas reacted with oxygen gas to yield water and chlorine gas. The mole ratio of the hydrogen chloride gas to water is
- A. 1 : 3.
 B. 2 : 1.
 C. 3 : 1.
 D. 4 : 1.

29. What number of moles of oxygen would exert a pressure of 10 atm at 320 K in an 8.2 dm³ cylinder?
 [R = 0.082 atm dm³ mol⁻¹ K⁻¹]
- A. 0.32
 B. 1.52
 C. 3.13
 D. 31.25
30. If 50 cm³ of a saturated solution of KNO₃ at 40 °C contained 5.05 g of the salt, its solubility at the same temperature would be [KNO₃ = 101]
- A. 1.0 mol dm⁻³.
 B. 1.5 mol dm⁻³.
 C. 2.0 mol dm⁻³.
 D. 5.0 mol dm⁻³.
31. Which of the following elements would displace copper from a solution of copper ions?
- A. Silver
 B. Gold
 C. Lead
 D. Mercury
32. What is the percentage composition of carbon in Ca(HCO₃)₂?
 [Ca = 40.0, O = 16.0, C = 12.0, H = 1.0]
- A. 22.2 %
 B. 14.8 %
 C. 7.4 %
 D. 3.7 %
33. Which of the following bond types is intermolecular?
- A. Covalent bond
 B. Hydrogen bond
 C. Ionic bond
 D. Metallic bond

Guideline For Chemistry Waec 2014 2015

**Randy E. Bennett, Linda Darling-
Hammond, Aneesha Badrinarayan**



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