

Guided Reading (Chapter 18 Section 2)

1. What do chemical bonds do?

Chemical bonds hold the atoms of molecules together and determine their chemical and physical properties.

2. Is a covalent bond more often found in atoms that are nonmetals without variable valence?

3. Why do some molecules accept H⁺ ions in the water solution? Hydrogen (water) bonded bonds, yes?

The water molecules have lone pairs of electrons that are able to form an H-bond.

4. Other groups than oxygen or nitrogen can form H-bonds, yes? Yes.

5. Is water liquid at room temperature because of hydrogen bonds?

6. What is a hydrogen bond?

The bonds result when one of the atoms of one molecule is bonded to two other atoms.

7. Why are the water gas molecules attracted to each other?

Because there are lone pairs of electrons on the oxygen.

8. How many electrons does the hydrogen atom have? Is hydrogen bond? Hydrogen has 1 valence electron.

9. What are the high-energy bond electrons of an atom called? Valence electrons.

Guided Reading Chapter 16 Section Covalent

Michal Rosen-Zvi



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