

(1.6-3)

From Appendix A.4,  $c_p = 4.02 \text{ kJ/kg} \cdot \text{K}$  at  $32.8^\circ\text{C}$   
Assume  $c_{p,m} = 4.02$  Note:  $(t-10)^\circ\text{C} = (t-10)\text{K}$

$$Q = 121\,300 \text{ kJ} = 454 \text{ kg} \left( 4.02 \frac{\text{kJ}}{\text{kg} \cdot \text{K}} \right) (t - 10) \text{ K}$$

Solving,  $(t-10) = 66.5 \text{ K} = 66.5^\circ\text{C}$   $t = 76.5^\circ\text{C}$

(1.6-4)

(a) A.3-9  $P = 30 \text{ psia}$  Liquid  $\text{H}_2\text{O}$

$$\Delta H = H_{240^\circ\text{F}} - H_{40^\circ\text{F}} = 208.44 - 8.02 = \boxed{200.42 \text{ Btu/lb}_m}$$

(b)  $P = 24.97 \text{ psia}$

$$H_{240^\circ\text{F}} (\text{vapor}) = 1160.7$$

$$H_{40^\circ\text{F}} (\text{liquid}) = 8.02$$

$$\Delta H = 1152.7 \text{ Btu/lb}_m$$

(c)  $P = 1.0 \text{ atm abs}$

$$H_{212^\circ\text{F}} (\text{vapor}) = 1150.5$$

$$H_{40^\circ\text{F}} (\text{liquid}) = 28.1$$

$$\Delta H = -1122.4 \text{ Btu/lb}_m$$

(d)  $P = 1.0 \text{ atm abs}$

$$H_{212^\circ\text{F}} (\text{vapor}) = 1150.5$$

$$H_{212^\circ\text{F}} (\text{liquid}) = 180.2$$

$$\Delta H = -970.3 \text{ Btu/lb}_m$$
$$(-2256.9 \text{ kJ/kg})$$

# Geankoplis Absorption Solution Manual

**Phillip C. Wankat**



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## **Table of Contents Geankoplis Absorption Solution Manual**

1. Understanding the eBook Geankoplis Absorption Solution Manual
  - The Rise of Digital Reading Geankoplis Absorption Solution Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Geankoplis Absorption Solution Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Geankoplis Absorption Solution Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Geankoplis Absorption Solution Manual
  - Personalized Recommendations
  - Geankoplis Absorption Solution Manual User Reviews and Ratings
  - Geankoplis Absorption Solution Manual and Bestseller Lists
5. Accessing Geankoplis Absorption Solution Manual Free and Paid eBooks
  - Geankoplis Absorption Solution Manual Public Domain eBooks
  - Geankoplis Absorption Solution Manual eBook Subscription Services
  - Geankoplis Absorption Solution Manual Budget-Friendly Options
6. Navigating Geankoplis Absorption Solution Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Geankoplis Absorption Solution Manual Compatibility with Devices
  - Geankoplis Absorption Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Geankoplis Absorption Solution Manual
  - Highlighting and Note-Taking Geankoplis Absorption Solution Manual
  - Interactive Elements Geankoplis Absorption Solution Manual

8. Staying Engaged with Geankoplis Absorption Solution Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Geankoplis Absorption Solution Manual
9. Balancing eBooks and Physical Books Geankoplis Absorption Solution Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Geankoplis Absorption Solution Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Geankoplis Absorption Solution Manual
  - Setting Reading Goals Geankoplis Absorption Solution Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Geankoplis Absorption Solution Manual
  - Fact-Checking eBook Content of Geankoplis Absorption Solution Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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