
LIMITING REACTANT LAB

MATERIALS: 250 mL beaker, 100 mL graduated cylinder, spoon or stirring rod, and a watch glass.

SAFETY PRECAUTIONS: Wear goggles during this experiment. Avoid skin contact with copper (II) chloride dihydrate. Wash your hands before leaving the laboratory.

PRELAB QUESTION:

Aluminum metal reacts with aqueous copper (II) chloride dihydrate ($\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$) to form copper metal, aqueous aluminum chloride, and water. Write the balanced equation for this reaction, including state symbols.

PROCEDURE:

DAY 1

1. Weigh out the copper (II) chloride dihydrate. Record the true mass of the hydrated compound used.
2. Using the chart below, select your group number and the approximate masses of aluminum metal and copper (II) chloride dihydrate you'll be using in this activity:

Groups	Mass of Al (g)	Mass of $\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$ (g)
1 & 7 & 13	1.00	2.00
2 & 8 & 14	0.50	2.00
3 & 9	0.50	4.00
4 & 10	0.50	6.00
5 & 11	0.25	5.00
6 & 12	0.25	4.00

3. Using a graduated cylinder, measure water (tap water is fine) and add it to your LABELED beaker containing the copper (II) chloride dihydrate. Stir until the salt dissolves completely. Keep adding water (maximum of 100 mL) until all salt is dissolved – record the amount of water used. Note the color of the solution.
4. Weigh out the appropriate mass of aluminum foil, record that mass, then crumple it up into a few small pieces (NOT TIGHTLY CRUMPLED) and add it to the copper (II) chloride dihydrate solution. Stir the solution gently and then allow it to sit undisturbed for 5-10 minutes. Write observations.

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S Ashworth



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Table of Contents Limiting Reactant Pogil Lab

1. Understanding the eBook Limiting Reactant Pogil Lab
 - The Rise of Digital Reading Limiting Reactant Pogil Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Pogil Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Pogil Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Pogil Lab
 - Personalized Recommendations
 - Limiting Reactant Pogil Lab User Reviews and Ratings
 - Limiting Reactant Pogil Lab and Bestseller Lists
5. Accessing Limiting Reactant Pogil Lab Free and Paid eBooks
 - Limiting Reactant Pogil Lab Public Domain eBooks
 - Limiting Reactant Pogil Lab eBook Subscription Services
 - Limiting Reactant Pogil Lab Budget-Friendly Options
6. Navigating Limiting Reactant Pogil Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Pogil Lab Compatibility with Devices
 - Limiting Reactant Pogil Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Pogil Lab
 - Highlighting and Note-Taking Limiting Reactant Pogil Lab
 - Interactive Elements Limiting Reactant Pogil Lab
8. Staying Engaged with Limiting Reactant Pogil Lab

-
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Pogil Lab
9. Balancing eBooks and Physical Books Limiting Reactant Pogil Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Pogil Lab
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactant Pogil Lab
 - Setting Reading Goals Limiting Reactant Pogil Lab
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactant Pogil Lab
 - Fact-Checking eBook Content of Limiting Reactant Pogil Lab
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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