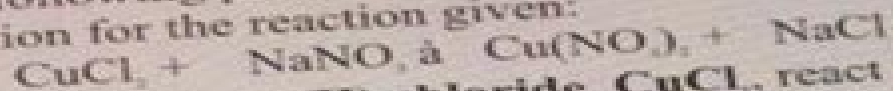


[Limiting Reactant and Percent Yield Worksheet]

Answer the following problems with complete solutions.

1. The equation for the reaction given:

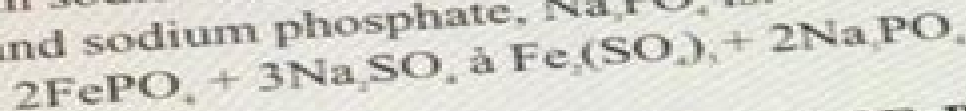


a) If 0.11 mol of copper (II) chloride, CuCl_2 , react with 0.24 mol of sodium nitrate, NaNO_3 , how many moles of sodium chloride, NaCl , can be formed?

b) What is the name of the limiting reagent?

c) If 11.3 grams of sodium chloride (actual yield) are formed in the reaction, what is the percent yield of this reaction?

2. The balanced equation for the reaction of iron (III) phosphate, FePO_4 , with sodium sulfate, Na_2SO_4 , to make iron (III) sulfate, $\text{Fe}_2(\text{SO}_4)_3$, and sodium phosphate, Na_3PO_4 , is:



a) If you perform this reaction with 0.17 mol of FePO_4 and an excess of Na_2SO_4 , how many moles of $\text{Fe}_2(\text{SO}_4)_3$ can you make?

b) If 18.5 grams of $\text{Fe}_2(\text{SO}_4)_3$ are actually made when you do this reaction, what is your percent yield?

Limiting Reactant And Percent Yield Lab Answers

**Christopher R. Hren, John T.
Moore, Peter J. Mikulecky**



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Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory. This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill. It begins with simple examples from everyday life, demonstrates the importance of balancing the equation, addresses the role of the mole in these computations, discusses different types of liquids, considers the role of significant figures, and culminates with the planning of syntheses. There are suggestions for further reading as well as practice problems and questions to ensure mastery. Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory. Addresses the necessity of balancing the reaction equation, the centrality of the mole in these calculations, and the role of significant figures in reporting the answer. Explains how to approach the calculations when using neat liquids or solutions. The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product. All of the problems in the book include detailed solutions with accompanying text to explain the answers, and ancillaries also include suggestions for further reading.

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