

### Limiting Reactant- Theoretical and Percentage Yields Key



If a reaction vessel contains 0.15 mol  $\text{KO}_2$  and 0.10 mol  $\text{H}_2\text{O}$ , what is the limiting reactant? How many moles of oxygen can be produced?

$$0.15 \text{ mol KO}_2 \times \frac{3 \text{ mol O}_2}{4 \text{ mol KO}_2} = 0.11 \text{ mol O}_2$$

$$0.10 \text{ mol H}_2\text{O} \times \frac{3 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}} = 0.15 \text{ mol O}_2$$

Limiting reactant:  $\text{KO}_2$

Maximum or theoretical yield = 0.11 mol  $\text{O}_2$



If a reaction vessel contains 10.0 g of sodium chloride and 12.0 g of sulfuric acid, what is the limiting reactant? What is the theoretical yield of hydrochloric acid? How many grams of excess reactant are left unconsumed?



$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{2 \text{ mol HCl}}{2 \text{ mol NaCl}} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 6.24 \text{ g HCl}$$

$$12.0 \text{ g H}_2\text{SO}_4 \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{2 \text{ mol HCl}}{1 \text{ mol H}_2\text{SO}_4} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.92 \text{ g HCl}$$

Limiting reactant:  $\text{NaCl}$

Maximum or theoretical yield = 6.24 g  $\text{HCl}$

$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{1 \text{ mol H}_2\text{SO}_4}{2 \text{ mol NaCl}} \times \frac{98.08 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = 8.39 \text{ g H}_2\text{SO}_4 \text{ required to consume all of the NaCl (Limiting reactant)}$$

$12.0 \text{ g H}_2\text{SO}_4 - 8.39 \text{ g H}_2\text{SO}_4 = 3.61 \text{ g}$  of excess  $\text{H}_2\text{SO}_4$  remains after reaction is complete

# Limiting Reagent Homework Answer Key

**G Thomas**



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