

# INORGANIC CHEMISTRY



electron  
valence  
in transition  
hydrogen

OR



energy level

is off  
2 with a  
even number

$d^1$   $d^9$   
 $n = 10^{10}$   
 $1.1 \times 10$   
 $1.66 \times 10$   
 $10^4$

**As**  
79  
 $3d^5 4s^1$   
1064.2  
29.76  
14.31  
 $+1, +3$

**Hg**  
201  
 $5d^{10} 6s^2$   
-18.8  
234.7  
13.546  
 $+1, +2$

hydrogen, for

## CHEMICAL BONDING

### IONIC BONDS

- Atoms with low ionization energy form cations (+), and atoms with high electron affinity form anions (-); attraction between opposite charges of cations and anions creates an ionic bond.

### THE LEWIS OCTET RULE: COVALENT BONDS

- Atoms other than hydrogen (stable with 2 valence electrons) will form bonds until they are surrounded by 8 valence electrons, and have achieved stable noble gas configuration.

The Lewis Dot Diagram  
Consists of one dot for each valence electron surrounding the symbol of the element.



### BOND DISSOCIATION ENERGIES: CALCULATING $\Delta H$

- $\Delta H = \sum \text{Bond energies of reactants} - \sum \text{Bond energies of products}$
- Example:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ , find the enthalpy of reaction,  $\Delta H$ .
- $\Delta H = 4(414.8) + 2(495.8) - 2(463.9) - 2(463.9) = -890.4 \text{ kJ/mol}$  (exothermic reaction)
- Reaction is exothermic = energy (heat) flows out of system
- Endothermic = energy (heat) flows into system

Bond	Energy (kJ/mol)
H - H	436.4
H - O	463
H - Cl	431.9
Cl - Cl	242.7
C - H	414
C - C	347

### PERCENT IONIC CHARACTER OF A CHEMICAL BOND

- Bonds with 50% or greater ionic character are ionic.
- Bonds with less than 50% ionic character are covalent.
- Transition from ionic to covalent bonding results in a reduction in electrical conductivity, melting point, and boiling point.



## ELECTRONIC CONFIGURATION OF THE ELEMENTS

### QUANTUM NUMBER & PROPERTIES

- Principal Quantum Number** ( $n = 1, 2, 3, 4, \dots$ )
  - Determines size of electron orbital (roughly proportional to  $n^2$ )
  - Angular Momentum Quantum Number ( $l = 0, 1, 2, \dots, (n-1)$ )
  - Determines shape of electron orbital
    - $l = 0$  Shape = spherical
    - $l = 1$  Shape = lobes
    - $l = 2$  Shape = lobes & donuts
- Magnetic Quantum Number** ( $m = 0$  for  $l = 0$ ;  $m = -l, -(l-1), \dots, 0, \dots, (l-1), l$  for  $l = 1, 2$ )
  - orientation of electron orbital
- Spin Quantum Number** ( $s = +\frac{1}{2}$  or  $-\frac{1}{2}$  for each  $l, m$ )
  - Determines spin of electron & (paired) (opposite) spins produce paramagnetism & (paired) spins produce diamagnetism

## INTRODUCTION

- The study of compounds (group of carbon, except for  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{CS}_2$ , and compounds containing  $\text{CN}^-$ ,  $\text{CO}_3^{2-}$ , and  $\text{HCO}_3^-$ )

### INORGANIC COMPOUNDS

Name	Formula	Name	Formula
Alumina	$\text{Al}_2\text{O}_3$	Glauber's Salt	$\text{Na}_2\text{SO}_4$
Ammonia	$\text{NH}_3$	Cyanate	$\text{CaNCN} \cdot 2\text{H}_2\text{O}$
Anhydrous	$\text{SO}_3$	Heavy Water	$\text{D}_2\text{O}$
Baking soda	$\text{NaHCO}_3$	Litharge	$\text{PbO}$
Bauxite	$\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O} \cdot \text{H}_2\text{O}$	Salt or Magnesia	$\text{MgCO}_3$
Calcite	$\text{CaCO}_3$	Devilstone	$\text{CaO}$
Carbon	$\text{H}_2\text{CO}_3$	Silica	$\text{SiO}_2$
Cyanide	$\text{Na}_2\text{CN}_2$	Sulphur Lime	$\text{CaSO}_4$
Dry Ice	$\text{CO}_2$ (solid)	Table Salt	$\text{NaCl}$
Gypsum	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$	Washing Soda	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$

### COMMON INORGANIC IONS

Ammonium ( $\text{NH}_4^+$ )	Mercuric ( $\text{Hg}_2^{2+}$ )	Nitride ( $\text{N}^{3-}$ )
Ammonium ( $\text{NH}_4^+$ )	Mercurous ( $\text{Hg}_2^{2+}$ )	Hydrogen Carbonate or Bicarbonate ( $\text{HCO}_3^-$ )
Barium ( $\text{Ba}^{2+}$ )	Potassium ( $\text{K}^+$ )	Hydrogen Phosphate ( $\text{HPO}_4^{2-}$ )
Calcium ( $\text{Ca}^{2+}$ )	Silver ( $\text{Ag}^+$ )	Hydrogen Sulfate, Bisulfate ( $\text{HSO}_4^-$ )
Cadmium ( $\text{Cd}^{2+}$ )	Sodium ( $\text{Na}^+$ )	Nitrate ( $\text{NO}_3^-$ )
Cerium ( $\text{Ce}^{3+}$ )	Strontium ( $\text{Sr}^{2+}$ )	Fluoride ( $\text{F}^-$ )
Chromium ( $\text{Cr}^{3+}$ )	Tin ( $\text{Sn}$ ), Stannous ( $\text{Sn}^{2+}$ )	Selenate ( $\text{SeO}_4^{2-}$ )
Cobalt ( $\text{Co}$ ), Cobaltous ( $\text{Co}^{2+}$ )	Zinc ( $\text{Zn}^{2+}$ )	Selenide ( $\text{Se}^{2-}$ )
Copper ( $\text{Cu}$ ), Cuprous ( $\text{Cu}^+$ )	Bromide ( $\text{Br}^-$ )	Nitrite ( $\text{NO}_2^-$ )
Copper ( $\text{Cu}$ ), Cupric ( $\text{Cu}^{2+}$ )	Carbonate ( $\text{CO}_3^{2-}$ )	Chloride ( $\text{Cl}^-$ )
Hydrogen ( $\text{H}^+$ )	Chloride ( $\text{ClO}_2^-$ )	Perchlorate ( $\text{ClO}_4^-$ )
Iron ( $\text{Fe}$ ), Ferrous ( $\text{Fe}^{2+}$ )	Chloride ( $\text{Cl}^-$ )	Formate ( $\text{HCO}_2^-$ )
Iron ( $\text{Fe}$ ), Ferric ( $\text{Fe}^{3+}$ )	Chromate ( $\text{CrO}_4^{2-}$ )	Phosphate ( $\text{PO}_4^{3-}$ )
Lead ( $\text{Pb}$ ), Plumbic ( $\text{Pb}^{2+}$ )	Cyanide ( $\text{CN}^-$ )	Sulfate ( $\text{SO}_4^{2-}$ )
Lithium ( $\text{Li}^+$ )	Dichromate ( $\text{Cr}_2\text{O}_7^{2-}$ )	Sulfide ( $\text{S}^{2-}$ )
Magnesium ( $\text{Mg}$ ), Magnesian	Ethylhydrogen Phosphate ( $\text{H}_2\text{PO}_4^-$ )	Sulfate ( $\text{SO}_4^{2-}$ )
Magnesium ( $\text{Mg}^{2+}$ )	Bromide ( $\text{Br}^-$ )	Thiosulfate ( $\text{S}_2\text{O}_3^{2-}$ )

### PAULI'S RULE

- When more than one orbital has the same energy, electrons occupy separate orbitals, and do so with opposite spins.
- Using 2p orbitals to demonstrate

### ORBITAL TYPES & SHAPES



### PAULI'S EXCLUSION PRINCIPLE

- No 2 electrons in an atom can have the same 4 quantum numbers. No more than 2 electrons can occupy any 1 orbital.

# Inorganic Elements Student Guide

**JG Myers**



## **Inorganic Elements Student Guide:**

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