

• N_2 is absorbed by CaC_2 at $1000^\circ C$
 $CaC_2 + N_2 \xrightarrow{1000^\circ C} CaCN_2 + C$

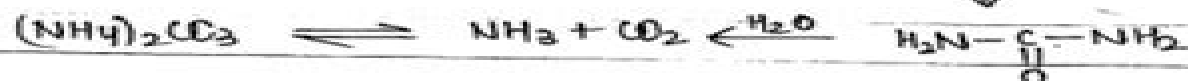
Ammonia (NH_3):

Preparation:

- ① Any NH_4^+ salt + Base $\rightarrow NH_3 \uparrow$
 $NH_4Cl + NaOH \rightarrow NH_3 \uparrow + NaCl$
- ② Any metal nitride + H_2O
 $Mg_3N_2 + H_2O \rightarrow Mg(OH)_2 + NH_3 \uparrow$
- ③ Heating of NH_4^+ salt having non/weak oxidising anion
 $(NH_4)_2CO_3 \xrightarrow{\Delta} NH_3 + H_2CO_3$
 $NH_4Cl \xrightarrow{\Delta} NH_3 + HCl$
 $NH_4CN \xrightarrow{\Delta} NH_3 + HCN$
- ④ $NO_2^- + \text{Fe} + NaOH \rightarrow NH_3 + NaNO_2$
 $\quad \quad \quad \uparrow$
 $\quad \quad \quad H_2$
- ⑤ $NO_3^- + Zn + NaOH \rightarrow NH_3 + Na_2ZnO_2$
- ⑥ Haber's process:
 $N_2 + 3H_2 \xrightarrow[\text{high pressure}]{Fe, Mo} 2NH_3 \quad \Delta H = -84$

Uses: ① Used in identification of transition metal cation.

② In preparation of nitrogen containing fertilizers.



Inorganic Chemistry Notes

ML Morrison



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1. Understanding the eBook Inorganic Chemistry Notes
 - The Rise of Digital Reading Inorganic Chemistry Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Inorganic Chemistry Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Inorganic Chemistry Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Inorganic Chemistry Notes
 - Personalized Recommendations
 - Inorganic Chemistry Notes User Reviews and Ratings
 - Inorganic Chemistry Notes and Bestseller Lists
5. Accessing Inorganic Chemistry Notes Free and Paid eBooks
 - Inorganic Chemistry Notes Public Domain eBooks

- Inorganic Chemistry Notes eBook Subscription Services
- Inorganic Chemistry Notes Budget-Friendly Options
- 6. Navigating Inorganic Chemistry Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Inorganic Chemistry Notes Compatibility with Devices
 - Inorganic Chemistry Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Inorganic Chemistry Notes
 - Highlighting and Note-Taking Inorganic Chemistry Notes
 - Interactive Elements Inorganic Chemistry Notes
- 8. Staying Engaged with Inorganic Chemistry Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Inorganic Chemistry Notes
- 9. Balancing eBooks and Physical Books Inorganic Chemistry Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Inorganic Chemistry Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Inorganic Chemistry Notes
 - Setting Reading Goals Inorganic Chemistry Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Inorganic Chemistry Notes
 - Fact-Checking eBook Content of Inorganic Chemistry Notes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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