

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^{-}$ to F	Li^{+}
F	7 ve	accepts $1e^{-}$ from Li	F^{-}

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$	Li	1 ve	donates $1e^{-}$	Li^{+1}
$\text{Li} + \text{Cl} \rightarrow [\text{Li}]^{+1} [\text{Cl}]^{-1}$	Cl	1 ve	accepts $1e^{-}$	Cl^{-1}
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$	Ca	2 ve	donates $2e^{-}$	Ca^{+2}
$\text{Ca} + \text{O} \rightarrow [\text{Ca}]^{+2} [\text{O}]^{2-}$	O	6 ve	accepts $2e^{-}$	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$	Be	2 ve	donates $2e^{-}$ to each F	Be^{+2}
$\text{Be} + \text{F} \rightarrow [\text{Be}]^{+2} [\text{F}]^{-1}$	F	7 ve	each F accepts $1e^{-}$ each	F^{-1}
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$	Mg	2 ve	donates $2e^{-}$ to S	Mg^{+2}
$\text{Mg} + \text{S} \rightarrow [\text{Mg}]^{+2} [\text{S}]^{2-}$	S	6 ve	accepts $2e^{-}$	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$	K	1 ve	donates $1e^{-}$	K^{+1}
$\text{K} + \text{F} \rightarrow [\text{K}]^{+1} [\text{F}]^{-1}$	F	7 ve	accepts $1e^{-}$	F^{-1}

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AW Chickering



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