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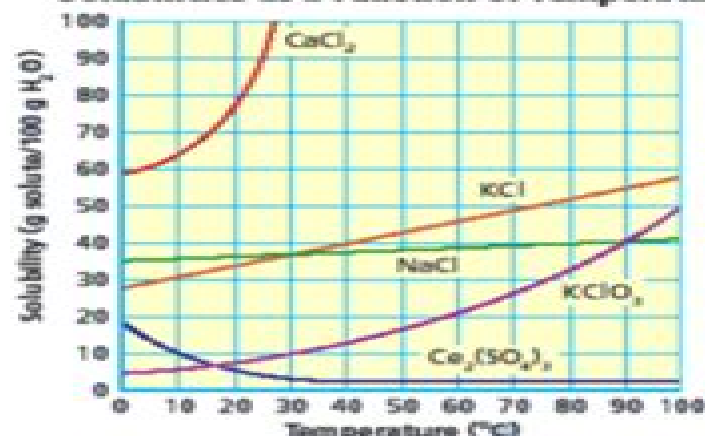
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## Solubility Curve Worksheet 2



Directions: Use the graph below to answer the following questions.

**Solubilities as a Function of Temperature**



### Types of Solutions

- Points that collectively make up the solubility curve (points **ON** the curve) represent \_\_\_\_\_ solutions.
- Points that are **BELOW** the curve represent \_\_\_\_\_ solutions.
- Points **ABOVE** the solubility curve represent \_\_\_\_\_ solutions, and the **difference** between the point above the curve and on the curve represents the amount of solute which will precipitate out.
- Say if the resulting solutions would be saturated, supersaturated or unsaturated.
  - 60 g of KCl at 70°C \_\_\_\_\_
  - 10 g of KClO<sub>3</sub> at 60°C \_\_\_\_\_
  - 70 g of CaCl<sub>2</sub> at 20°C \_\_\_\_\_

1. a. What is the solubility of calcium chloride (CaCl<sub>2</sub>) at 5°C? \_\_\_\_\_  
 b. What is the solubility of calcium chloride at 25°C? \_\_\_\_\_  
 c. What is happening to the solubility and temperature for CaCl<sub>2</sub>? Solubility is \_\_\_\_\_ as temperature is \_\_\_\_\_.
2. a. What is the solubility of cerium sulfate - Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> at 10°C? \_\_\_\_\_  
 b. What is the solubility of cerium sulfate at 50°C? \_\_\_\_\_  
 c. What is happening to the solubility and temperature for Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>? Solubility is \_\_\_\_\_ as temperature is \_\_\_\_\_.
3. Which substance on the graph is **least** soluble at 10°C? \_\_\_\_\_
4. Which substance on the graph shows the **least** change in solubility from 0°C to 100°C? \_\_\_\_\_
5. Which solution is more concentrated:
  - a) At 10°C, a saturated solution of NaCl or a saturated solution of CaCl<sub>2</sub> (circle one)
  - b) At 55°C, a saturated solution of KClO<sub>3</sub> or a saturated solution of Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> (circle one)
6. What is the mass of cerium sulfate that will dissolve in 50 g of water at 10 degrees Celsius?
7. What is the mass of potassium chloride that will dissolve in 250 g of water at 70 degrees Celsius?
8. At 90 degrees Celsius, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? How do you know?
9. A saturated solution of potassium chlorate is dissolved in 100g of water. If the saturated solution is cooled from 90 degrees Celsius to 60 degrees Celsius, how many grams would crystallize out?

# Graphing Solubility Curve Practice Problems

**Katrin Zwirgmaier**



## **Graphing Solubility Curve Practice Problems:**

This book delves into Graphing Solubility Curve Practice Problems. Graphing Solubility Curve Practice Problems is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Graphing Solubility Curve Practice Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
    - Chapter 1: Introduction to Graphing Solubility Curve Practice Problems
    - Chapter 2: Essential Elements of Graphing Solubility Curve Practice Problems
    - Chapter 3: Graphing Solubility Curve Practice Problems in Everyday Life
    - Chapter 4: Graphing Solubility Curve Practice Problems in Specific Contexts
    - Chapter 5: Conclusion
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  3. In chapter 2, the author will delve into the foundational concepts of Graphing Solubility Curve Practice Problems. The second chapter will elucidate the essential principles that need to be understood to grasp Graphing Solubility Curve Practice Problems in its entirety.
  4. In chapter 3, this book will examine the practical applications of Graphing Solubility Curve Practice Problems in daily life. The third chapter will showcase real-world examples of how Graphing Solubility Curve Practice Problems can be effectively utilized in everyday scenarios.
  5. In chapter 4, the author will scrutinize the relevance of Graphing Solubility Curve Practice Problems in specific contexts. This chapter will explore how Graphing Solubility Curve Practice Problems is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, the author will draw a conclusion about Graphing Solubility Curve Practice Problems. The final chapter will summarize the key points that have been discussed throughout the book.
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