

$$\frac{214.2 \text{ g OsF}_6}{0.0673 \text{ L soln}} \times \frac{1 \text{ mol OsF}_6}{247.23 \text{ g OsF}_6} = 12.9 \text{ M OsF}_6$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

7. A chalice contains 36.45 grams ammonium chlorite in 2.36 liters of solution - calculate the molarity.

$$\frac{36.45 \text{ g NH}_4\text{ClO}_2}{2.36 \text{ L soln}} \times \frac{1 \text{ mol NH}_4\text{ClO}_2}{85.50 \text{ g NH}_4\text{ClO}_2} = 0.181 \text{ M NH}_4\text{ClO}_2$$

8. What is the molarity of a solution that contains 14.92 grams magnesium oxalate in 3.65 ml of solution?

$$\frac{14.92 \text{ g MgC}_2\text{O}_4}{0.00365 \text{ L soln}} \times \frac{1 \text{ mol MgC}_2\text{O}_4}{112.32 \text{ g MgC}_2\text{O}_4} = 36.4 \text{ M MgC}_2\text{O}_4$$

9. What mass of lithium phosphate would you mass to make 2.5 liter of 1.06 M lithium phosphate solution?

$$2.5 \text{ L soln} \times \frac{1.06 \text{ mol Li}_3\text{PO}_4}{1 \text{ L soln}} \times \frac{115.79 \text{ g Li}_3\text{PO}_4}{1 \text{ mol Li}_3\text{PO}_4} = 310 \text{ g Li}_3\text{PO}_4$$

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