

Chemistry Practice Test 1 Solution

1. Select the best word from the list to complete each of the following statements. (14 marks)

elements, gas, groups, formula, ion, neutralisation, negative, combustion,
decomposition, metals, neutrons, hydroxide, carbonate, positive

- a. CO_3^{2-} is called the carbonate ion.
- b. The formula for Sodium hydroxide is NaOH .
- c. Iron, copper and lead are classified as metals.
- d. An acid reacting with a carbonate produces bubbles of carbon dioxide gas.
- e. An atom that has gained or lost electrons is called an ion.
- f. One substance splitting into two or more substances is called decomposition.
- g. Protons and neutrons are found in the nucleus of atoms.
- h. OH^- is called the hydroxide ion.
- i. Columns in the periodic table are called groups.
- j. The reaction between an acid and a base is called neutralisation.
- k. The periodic table contains only elements.
- l. The word combustion means burning.
- m. Protons have a positive charge.
- n. Electrons have a negative charge.

2. Using the Periodic Table, complete the table below. (4 marks)

Element Name	Element Symbol	Atomic number	Number of protons	Number of electrons
Lithium	Li	3	3	3

3. Complete the following table. (6 marks)

Reaction	Reactants	Products
Precipitation	Two soluble salt solutions	A soluble salt solution and an insoluble salt
Neutralisation	Acid and base	A salt and water
Acid on Carbonate	Acid and carbonate	A salt, water and carbon dioxide gas
Acid on Metal	Acid and metal	A salt and hydrogen gas

4. Identify the mystery substance (P,Q,R,S,T,U) for each reaction. (6 marks)

Reaction	Mystery substance
Neutralisation: Hydrochloric acid is added to P and sodium chloride and water are produced.	P: Sodium hydroxide
Decomposition: Electricity is passed through water and Q and oxygen gases are produced.	Q: Hydrogen
Combustion: Wood is heated in air and a gas R and water are produced.	R: Carbon dioxide
Precipitation: Aqueous silver nitrate is added to aqueous sodium chloride and S, a white solid, is observed to have formed at the bottom of the reaction vessel.	S: Silver chloride
Acid on a carbonate: Hydrobromic acid is added to calcium carbonate and water, calcium bromide and a gas T are produced. The gas turns lime water milky.	T: Carbon dioxide
MASH: Magnesium metal is placed in acid and a gas U is formed.	U: Hydrogen

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Table of Contents Gradpoint Chemistry Review Test

1. Understanding the eBook Gradpoint Chemistry Review Test
 - The Rise of Digital Reading Gradpoint Chemistry Review Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Gradpoint Chemistry Review Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gradpoint Chemistry Review Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gradpoint Chemistry Review Test
 - Personalized Recommendations
 - Gradpoint Chemistry Review Test User Reviews and Ratings
 - Gradpoint Chemistry Review Test and Bestseller Lists

5. Accessing Gradpoint Chemistry Review Test Free and Paid eBooks
 - Gradpoint Chemistry Review Test Public Domain eBooks
 - Gradpoint Chemistry Review Test eBook Subscription Services
 - Gradpoint Chemistry Review Test Budget-Friendly Options
6. Navigating Gradpoint Chemistry Review Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Gradpoint Chemistry Review Test Compatibility with Devices
 - Gradpoint Chemistry Review Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gradpoint Chemistry Review Test
 - Highlighting and Note-Taking Gradpoint Chemistry Review Test
 - Interactive Elements Gradpoint Chemistry Review Test
8. Staying Engaged with Gradpoint Chemistry Review Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gradpoint Chemistry Review Test
9. Balancing eBooks and Physical Books Gradpoint Chemistry Review Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gradpoint Chemistry Review Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gradpoint Chemistry Review Test
 - Setting Reading Goals Gradpoint Chemistry Review Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gradpoint Chemistry Review Test
 - Fact-Checking eBook Content of Gradpoint Chemistry Review Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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