
ACS Gen Chem 1 Exam

Ion - A species of an element in which the number of electrons does not equal the number of protons.

Isotope - Each isotope of an atom has a different number of neutrons.

Isoelectric - Two atoms with the same charge.

Isotopic - Two atoms with the same number of neutrons.

Transition Metal Orbitals - The 4s orbital is filled before the 3d orbital, but listed after.

Paramagnetic - Atoms that possess a permanent magnetic charge due to the presence of unpaired electrons.

Electropositive - Tending to lose electrons and form positive ions.

Tetrahedral Bond Angle - 109.5°

Resonance Structures - Two equally correct arrangements of one molecule's electrons; generally the molecule exists as an average of the two.

Polar Molecule - A molecule that is mostly positive on one side and mostly negative on the other.

sp^3 Orbital - An orbital composed of one s orbital and three p orbitals; can hold eight total electrons. Tetrahedral.

Trigonal Planar - The way that three objects will arrange themselves around a central atom. Non-polar.

Free Radicals - Compounds with an odd number of electrons in their Lewis Structure. Eg. NO_2

Incomplete Octets - Occur in elements that can form stable bonds with less than eight electrons. Eg. BF_3

Expanded Octets - Occur in the third row of the periodic table and beyond in elements that can form stable bonds with more than eight electrons. Eg. XeF_2

Hydrogen Bonding - When hydrogen is attracted to electronegative atoms.

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Michael Hoelscher



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