

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



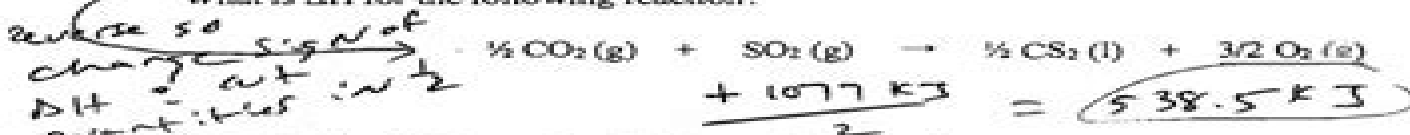
$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?



3. Following is the equation for the combustion of acetone:



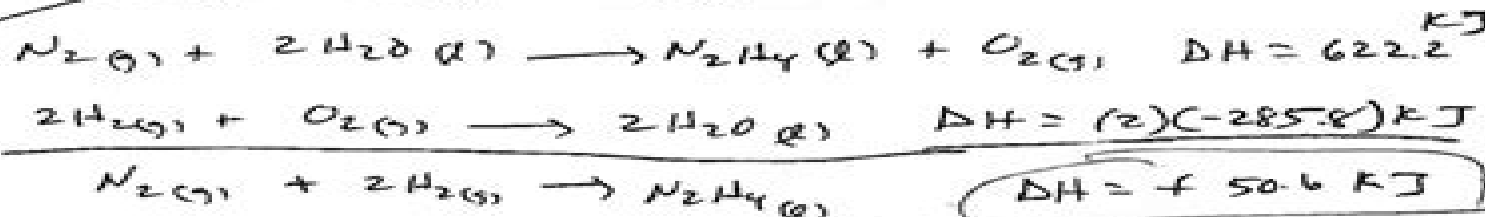
How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



General Chemistry Practice Problems With Answers

Susan E. Kegley, Allan R. Pinhas



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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