

**Chem 101 Practice Test: Thermodynamics 3.5 – 3.6, 18.1****Equations**

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta H_{\text{sys}} - T\Delta S_{\text{sys}}$$

$$\Delta G = \Delta H_{\text{sys}}$$

$$\Delta H_{\text{sys}} = \sum \text{products} - \sum \text{reactants}$$

$$\Delta S = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

**Properties of water**

$$\Delta_{\text{vap}} H^\circ = 40.7 \text{ kJ/mol}$$

$$\Delta_{\text{fus}} H^\circ = 6.01 \text{ kJ/mol}$$

$$\Delta_{\text{sub}} H^\circ = 47.05 \text{ kJ/mol}$$

$$\Delta_{\text{vap}} S^\circ = 109 \text{ J/mol}\cdot\text{K}$$

$$\Delta_{\text{fus}} S^\circ = 22 \text{ J/mol}\cdot\text{K}$$

**Multiple choice**

Multiple choice and calculation questions are worth 5 points. The calculations will be graded.

(1) spontaneous

(2) entropy

(3) phase

(4) enthalpy

(5) heat capacity

(6) specific heat

(7) boiling point

(8) heat of vaporization

(9) heat of fusion

1. spontaneity of a process is determined by the sign of  $\Delta G$ .
2.  $\Delta G$  is the difference between the free energy of the products and the free energy of the reactants.
3. spontaneity of a process is determined by the sign of  $\Delta H$ .
4.  $\Delta H$  is the difference between the enthalpy of the products and the enthalpy of the reactants.
5.  $\Delta S$  is the difference between the entropy of the products and the entropy of the reactants.
6.  $\Delta S$  is the difference between the entropy of the products and the entropy of the reactants.
7.  $\Delta S$  is the difference between the entropy of the products and the entropy of the reactants.

**Short answer questions**

1. Which of the following is a spontaneous process?
- a. A reaction with a positive  $\Delta G$
- b. A reaction with a negative  $\Delta G$
- c. A reaction with a positive  $\Delta H$
- d. A reaction with a negative  $\Delta H$
2. Which of the following is a spontaneous process?
- a. A reaction with a positive  $\Delta G$
- b. A reaction with a negative  $\Delta G$
- c. A reaction with a positive  $\Delta H$
- d. A reaction with a negative  $\Delta H$
3. Which of the following is a spontaneous process?
- a. A reaction with a positive  $\Delta G$
- b. A reaction with a negative  $\Delta G$
- c. A reaction with a positive  $\Delta H$
- d. A reaction with a negative  $\Delta H$

# Genchem Practice Test Thermochemistry

**Katrin Zwirgmaier**



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