



<b>Abbreviations, annotations and conventions used in the Mark Scheme</b>	/ = alternative and acceptable answers for the same marking point : = separates marking points NOT = answers which are not worthy of credit () = words which are not essential to gain credit <u>      </u> = (underlining) key words which <u>must</u> be used to gain credit ecf = error carried forward AW = alternative wording ora = or reverse argument
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Question	Expected answers	Marks
1 (a)	  ✓ 1, 2 or 3 p orbitals are OK	[2]
(b)	d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓	[3]
(c) (i) (ii)	 2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓	[2]
(d) (i)	Energy change when each atom in 1 mole ✓ of gaseous atoms ✓ loses an electron ✓ (to form 1 mole of gaseous 1+ ions).	[3]
(ii)	$O^{2+}(g) \longrightarrow O^{3+}(g) + e^{-}$ ✓✓ 1 mark for correct species; 1 mark for state symbols No charge required on electron. Ignore (g) on e	[2]

# F321 June 2013 Past Paper

**Peter Cann, Peter Hughes**



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## **Table of Contents F321 June 2013 Past Paper**

1. Understanding the eBook F321 June 2013 Past Paper
  - The Rise of Digital Reading F321 June 2013 Past Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying F321 June 2013 Past Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an F321 June 2013 Past Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from F321 June 2013 Past Paper
  - Personalized Recommendations
  - F321 June 2013 Past Paper User Reviews and Ratings

- F321 June 2013 Past Paper and Bestseller Lists
- 5. Accessing F321 June 2013 Past Paper Free and Paid eBooks
  - F321 June 2013 Past Paper Public Domain eBooks
  - F321 June 2013 Past Paper eBook Subscription Services
  - F321 June 2013 Past Paper Budget-Friendly Options
- 6. Navigating F321 June 2013 Past Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - F321 June 2013 Past Paper Compatibility with Devices
  - F321 June 2013 Past Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of F321 June 2013 Past Paper
  - Highlighting and Note-Taking F321 June 2013 Past Paper
  - Interactive Elements F321 June 2013 Past Paper
- 8. Staying Engaged with F321 June 2013 Past Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers F321 June 2013 Past Paper
- 9. Balancing eBooks and Physical Books F321 June 2013 Past Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection F321 June 2013 Past Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine F321 June 2013 Past Paper
  - Setting Reading Goals F321 June 2013 Past Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of F321 June 2013 Past Paper
  - Fact-Checking eBook Content of F321 June 2013 Past Paper
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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