Quiz)

Name ANSWER WEYERS

Date

Co HIDO

Assessment

Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- 1. Why is freezing-point depression a colligative property?
 - a. It is inversely proportional to the number of particles in a solution.
 - b. It is directly proportional to the number of particles in a solution.
 - c. It depends on the properties of an electrolyte in a solvent.
 - d. None of the above
- Compared with a 0.01 M C₀H₁₂O₀ solution, a 0.01 M KCl solution has
 - a. the same freezing-point depression.
 - b. about twice the freezing-point depression.
 - c. the same freezing-point elevation.
 - d. about six times the freezing-point elevation.
- 3. Compared with a 1.00 M NaI solution, a 1.00 M Na₂SO₄ solution has
 - a. the same boiling-point elevation.
 - b. about twice the boiling-point elevation.
 - c. a boiling-point elevation about two-thirds as high.
 - d. a boiling-point elevation about 1.5 times as high.
- 4. When a nonvolatile solute dissolves in a solvent, the vapor pressure of the solvent
 - a. increases.
 - b. decreases.
 - c. stays the same.
 - changes depending on the solvent used.
- 5. Which of the following compounds would be most effective in lowering the melting point of ice on roads?
 - a. CaCl.,
 - b. NaCl.
 - c. K₀PO₄
 - d. K₂SO₄

Nacl Hio Nat

(9) + C/F)

VAPOR PRESSURE LOWERING (Colligative

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Modern Chemistry

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Math Ch 13 Quiz 2 Answer Key

Robert R. Redfield

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