

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the **best** (but not all) of the information for this chapter is that.

Focus on this content, but make sure to review other topics: reactions, bonding, molecules, etc.

If your school has discussed and **NOTED** this, you should at least be able to P&M that text.

**** Best items will be recall, examples, and/or application of this content. ****

OUTLINE:

- Collapsing into periodicity (understanding/chemistry concepts: **PT** & **Q**)
- Communication/chemistry concepts: **to** quantum and periodic **PT** & **Q**
- 5.1: Describe the organization of the modern Periodic table: distinguish between periods and groups **PT** & **Q**
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical, correlation strategy, and atomic size) of that elements **PT** & **Q**
- 5.3: Identify common compounds, metalloids, and noble gases: identify the representative groups and/or main groups (groups: 1, 2, 13, 14, 15, 16, 17, 18) **PT** & **Q**
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity **PT** & **Q**

5.1: HISTORY OF THE PERIODIC TABLE

• Introductory

- **periodic law** = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
- **periodic table** = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- **classification** = a method of the elements, such as the number of electrons, relative atomic numbers, range from 11 (sodium) to 118 (Xenon)
- **classification** = one of the ways of classifying elements that elements from 1 (hydrogen) through 108 (unlabeled) (element number 109) are the periodic table

• Chapter Highlights

- The periodic law states that the physical and chemical properties of the elements, and periodic functions of their atomic numbers.
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- **atomic shell** = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- **electron configuration** = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- **transition elements** = one of the metals that can use the inner shell electrons using the outer shell to form
- **transition elements** = an element in the d-block or f-block of the periodic table
- **halogens** = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts.

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