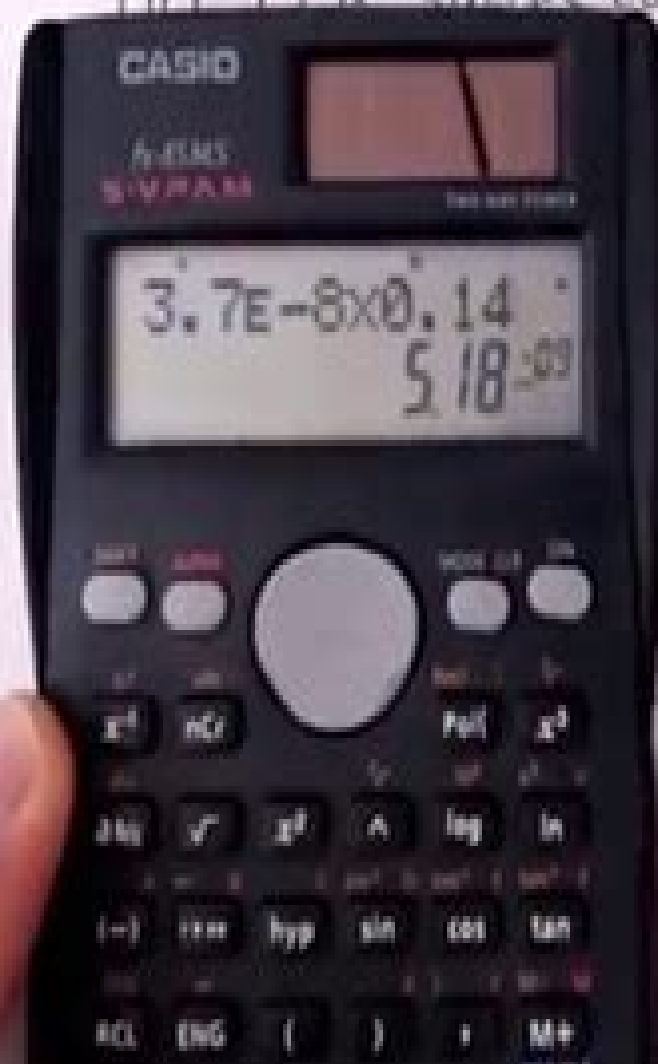


- (a)  $0.14 \text{ mol dm}^{-3}$  solutions of hydrochloric acid,  $\text{HCl}$ , and chloroic acid,  $\text{HClO}$  ( $K_a = 1.7 \times 10^{-8}$ ), have different pH values.

Explain why the pH values are different and calculate the pH of  $0.14 \text{ mol dm}^{-3}$  solutions of  $\text{HCl}$  and  $\text{HClO}$  to two decimal places.

Show any working in calculations.

$\text{HCl}$  is a strong acid and  $\text{HClO}$  is



$$\rightarrow -\log 0.14 = 0.85.$$

$$7 \times 10^{-8}$$

$$[\text{H}^+]^2 = 3.7 \times 10^{-8} \times 0.14$$

$$= 5.18 \times 10^{-9}$$

# F325 June 2013 Past Paper Pack

**TD Snyder**



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