

Empirical Formula

Objectives

- To determine the empirical formula of a compound by use of a combination reaction.
- To practice techniques for obtaining and recording precise mass values.
- To learn safe practices of Bunsen burner use.

Introduction

The empirical formula of a compound is the simplest (smallest) whole-number ratio of moles of elements in the compound. Determination of the empirical formula requires three steps:

- Obtain the actual mass of each element in the compound
- Calculate the moles of each element from its actual mass and molar mass
- Express the ratio of moles of each element as small whole numbers

Example

Elemental analysis of a sample of an unknown crystalline solid reveals the presence of 5.52 g of sodium and 8.48 g of chlorine as the only components. The moles of each element are calculated:

$$\text{Moles of sodium: } 5.52 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.99 \text{ g Na}} = 0.240 \text{ mol Na}$$

$$\text{Moles of chlorine: } 8.48 \text{ g Cl} \times \frac{1 \text{ mol Cl}}{35.45 \text{ g Cl}} = 0.239 \text{ mol Cl}$$

The mole ratio of sodium to chlorine in the compound is 0.240 : 0.239. Expressing the mole ratio as the simplest whole-number mole ratio reveals 1 mol Na : 1 mol Cl. The empirical formula is **NaCl**, and the compound is identified as *sodium chloride* (table salt).

The empirical formula of a *metal oxide* can be determined by carrying out a **combination reaction** at high temperature in air. The mass of the metal before the reaction, and the total mass of the metal oxide after the reaction are measured. From the difference between the masses, the mass of oxygen in the metal oxide is determined. The moles of each element are calculated and the mole ratio of the two elements yields the empirical formula for the metal oxide.

Example

A 5.90 g sample of titanium metal combines with oxygen in the air to form 9.84 g of titanium oxide. The empirical formula of the titanium oxide can be determined as follows:

$$\text{Moles of titanium: } 5.90 \text{ g Ti} \times \frac{1 \text{ mol Ti}}{47.87 \text{ g Ti}} = 0.123 \text{ mol Ti}$$

$$\text{Mass of oxygen: } 9.84 \text{ g compound} - 5.90 \text{ g Ti} = 3.94 \text{ g O}$$

$$\text{Moles of oxygen: } 3.94 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.246 \text{ mol O}$$

The mole ratio of titanium to oxygen in the compound is 0.123 : 0.246. Expressing the mole ratio as the simplest whole-number mole ratio reveals 1 mol Ti : 2 mol O. The empirical formula is **TiO₂**, and the compound is identified as *titanium(IV) oxide*.

Empirical Formula Of A Compound Lab Report

RC Schank

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