

Let the relative atomic mass of element Z be z .

Element	Z	Bromine, Br
Mass (g)	2.07	$3.67 - 2.07 = 1.60$
Number of moles of atoms	$\frac{2.07}{z}$	$\frac{1.60}{80} = 0.02$

Based on the empirical formula ZBr_2 , the ratio of atoms of Z : Br is 1 : 2.

Therefore, $\frac{2.07}{z} : 0.02 = 1 : 2$

$$\frac{2.07}{z} \div 0.02 = 1 \div 2$$

$$\frac{2.07}{0.02z} = \frac{1}{2}$$

$$z = \frac{2 \times 2.07}{0.02}$$

$$= 207$$

The relative atomic mass of element Z is 207.

Empirical Formula For Copper Oxide

Lingjun Ying

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Empirical Formula For Copper Oxide:

Unveiling the Magic of Words: A Report on "**Empirical Formula For Copper Oxide**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Empirical Formula For Copper Oxide**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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